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342g of 20% by mass of Ba(OH)_2 solution specific gravity 1.057 is reacted with 1200 ml of 2M HNO_3 . If the final density of solution is same as pure water, then molarity of the iron in resulting solution which decides the nature of the above solution is

- 1.025M
- 0.5M
- 0.888M
- none of

web what is the formula for converting molarity to molality molality is equal to the product of molarity divided by the density of the solution at that temperature $m = \frac{M}{d}$ which is more important molality or molarity both molality and molarity are concentration units and they are used interchangeably

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web 1 molarity $m = \frac{\text{mass}}{\text{volume}}$ mass molar mass $\times 0.4500 \text{ L} = 0.825 \text{ g}$ 141.9579 g/mol $\times 0.0129 \text{ m}^3 = 0.001825 \text{ g}$ 141.9579 g/mol 0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 m³ mole fraction $n_2 = \frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ $n_1 = \frac{450.0 \text{ g}}{18.015 \text{ g/mol}} = 24.97918401 \text{ mol}$

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