

Name: _____

Balancing Chemical Equations

Directions: Balance each of the following chemical equations.



- 1) ___ H_3PO_4 + ___ KOH $\rightarrow$ ___ K_3PO_4 + ___ H_2O
- 2) ___ K + ___ B_2O_3 $\rightarrow$ ___ K_2O + ___ B
- 3) ___ HCl + ___ NaOH $\rightarrow$ ___ NaCl + ___ H_2O
- 4) ___ Na + ___ NaNO_3 $\rightarrow$ ___ Na_2O + ___ N_2
- 5) ___ C + ___ S_8 $\rightarrow$ ___ CS_2
- 6) ___ Na + ___ O_2 $\rightarrow$ ___ Na_2O
- 7) ___ N_2 + ___ O_2 $\rightarrow$ ___ N_2O_5
- 8) ___ H_3PO_4 + ___ $\text{Mg}(\text{OH})_2$ $\rightarrow$ ___ $\text{Mg}_3(\text{PO}_4)_2$ + ___ H_2O
- 9) ___ NaOH + ___ H_2CO_3 $\rightarrow$ ___ Na_2CO_3 + ___ H_2O
- 10) ___ KOH + ___ HBr $\rightarrow$ ___ KBr + ___ H_2O
- 11) ___ Na + ___ O_2 $\rightarrow$ ___ Na_2O
- 12) ___ $\text{Al}(\text{OH})_3$ + ___ H_2CO_3 $\rightarrow$ ___ $\text{Al}_2(\text{CO}_3)_3$ + ___ H_2O
- 13) ___ Al + ___ S_8 $\rightarrow$ ___ Al_2S_3
- 14) ___ Cs + ___ N_2 $\rightarrow$ ___ Cs_3N
- 15) ___ Mg + ___ Cl_2 $\rightarrow$ ___ MgCl_2
- 16) ___ Rb + ___ RbNO_3 $\rightarrow$ ___ Rb_2O + ___ N_2
- 17) ___ C_6H_6 + ___ O_2 $\rightarrow$ ___ CO_2 + ___ H_2O
- 18) ___ N_2 + ___ H_2 $\rightarrow$ ___ NH_3

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Glencoe Science McGraw-Hill Staff,2001-06 *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science* ,2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction Resources in Education ,1998 Mathematical Methods in Chemical and Biological Engineering Binay Kanti Dutta,2016-11-03 Mathematical Methods in Chemical and Biological Engineering describes basic to moderately advanced mathematical techniques useful for shaping the model based analysis of chemical and biological engineering systems Covering an ideal balance of basic mathematical principles and applications to physico chemical problems this book presents examples drawn from recent scientific and technical literature on chemical engineering biological and biomedical engineering food processing and a variety of diffusional problems to demonstrate the real world value of the mathematical methods Emphasis is placed on the background and physical understanding of the problems to prepare students for future challenging and innovative applications *The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB Te HS&T a Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004-02 Chemical Interactions ,2005 *Science Spectrum* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 **Basic Skills Wkshts Sci Spectrum 2001** Holt Rinehart & Winston, 2000-03 **Te HS&T 2007 Shrt Crs M** Holt Rinehart & Winston, 2007

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