- 2 -

- Which of the following processes provide evidence of the particulate nature of matter?
  - I. Capillarity
  - II. Diffusion
  - III. Filtration
  - IV. Osmosis
  - (A) I and III only
  - (B) I and IV only
  - (C) II and III only
  - (D) II and IV only
- 'Mass number' is the number of
  - (A) neutrons plus protons
  - (B) neutrons minus protons
  - (C) electrons plus neutrons
  - (D) electrons plus protons
- 3. Which of the following elements has 7 electrons in its outer shell?
  - (A) Hydrogen
  - (B) Oxygen
  - (C) Nitrogen
  - (D) Chlorine
- An atom that has 20 electrons when it ionizes has the same electronic configuration as
  - (A) argon
  - (B) calcium
  - (C) neon
  - (D) sodium
- <sup>34</sup>Na and <sup>25</sup>Na are isotopes because they
  - (A) conduct electricity
  - (B) react vigorously with water
  - C) have the same number of protons
  - (D) have the same relative atomic mass

- Sulfur and oxygen are in the same group of the periodic table because
  - (A) they can react with each other
  - (B) they can form covalent compounds
  - (C) they have the same number of electrons in their outer shell
  - (D) the atomic number of sulfur is 16 and the relative atomic mass of oxygen is 16

<u>Item 7</u> refers to the following quantities of atoms.

- 12 g of carbon
- II. 16 g of sulfur
- III. 23 g of sodium
- IV. 12 g of magnesium

(Relative mass number: C = 12; S = 32; Mg = 24; Na = 23)

- 7. Which of the following pairs represents equal quantities of atoms?
  - (A) I and II only
  - (B) I and IV only
  - (C) II and III only
  - (D) II and IV only
- An ion with a single negative charge may be converted into a neutral atom by
  - (A) losing a neutron
  - (B) gaining a proton
  - (C) losing an electron
  - (D) gaining an electron

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  - The Rise of Digital Reading Gcse Physics June 2015 Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Gcse Physics June 2015 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gcse Physics June 2015 Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Gcse Physics June 2015 Answers

- Personalized Recommendations
- Gcse Physics June 2015 Answers User Reviews and Ratings
- Gcse Physics June 2015 Answers and Bestseller Lists
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  - Gcse Physics June 2015 Answers Budget-Friendly Options
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  - Gcse Physics June 2015 Answers Enhanced eBook Features
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  - Participating in Virtual Book Clubs
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  - Dealing with Digital Eye Strain
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  - Managing Screen Time
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  - Setting Reading Goals Gose Physics June 2015 Answers
  - Carving Out Dedicated Reading Time
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- Fact-Checking eBook Content of Gcse Physics June 2015 Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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