

# Gas Laws HOMEWORK

## Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 750 mmHg to atm

b. 550 mmHg to atm

c. 2.5 atm to torr

d.  $2.1 \times 10^3$  Pa to atm

e. 30 Pa to atm

$$750 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.9868 \text{ atm}$$

$$2.5 \text{ atm} \times (760 \text{ mmHg} / 1 \text{ atm}) = 1900 \text{ mmHg}$$

$$2013 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.01986 \text{ atm}$$

$$1700 \text{ mmHg} \times (133 \text{ Pa} / 1 \text{ mmHg}) = 227100 \text{ Pa}$$

$$2.2 \times 10^3 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.0217 \text{ atm}$$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 750 mm.

$$750 \text{ mm} \times (1013 \text{ Pa} / 760 \text{ mm}) = 100387 \text{ Pa} \rightarrow 100.4 \text{ kPa}$$

4. Calculate the pressure of a sample of gas (in mmHg) inside an open-end manometer if the mercury column is 35 mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 760 mm. Draw a picture to illustrate your answer.

$$760 \text{ mm} + (1700 \text{ mmHg} / 1 \text{ mm}) = 1744 \text{ mmHg}$$

\*22 The number is 17 mm Hg above the atmospheric pressure, then it's 17 mmHg above 760 mmHg.

\*Draw picture showing a manometer with the height of the mercury 35 mm above the atmospheric level of 760 mm.

$$760 \text{ mmHg} + 35 = 795 \text{ mmHg}$$

5. If a gas at a pressure of 150 mm in a 1.0 L container has its pressure increased to 400 mm, what is the new volume of the container? (Assume temperature is constant).

$$P_1 V_1 = P_2 V_2$$

$$(150 \text{ mm} \times 1.0 \text{ L}) = 400 \text{ mm} \times V_2$$

$$V_2 = 0.375 \text{ L} \rightarrow 0.38 \text{ L}$$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.5 atm is increased to 8.0 L?

$$P_1 V_1 = P_2 V_2$$

$$(2.5 \text{ atm} \times 5.00 \text{ L}) = P_2 \times 8.0 \text{ L}$$

$$P_2 = 1.56 \text{ atm}$$

7. When 0.5 L of nitrogen and 0.5 L of sulfur dioxide at  $0^\circ\text{C}$  and 1.0 atm, are pumped into an evacuated 1.0 L rigid container, what is the final pressure in the container at  $0^\circ\text{C}$ ? *Since temperature doesn't change, you can leave it out of the equation.*

So final pressure of Nitrogen:

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 0.5 \text{ atm}$$

So final pressure of Sulfur dioxide:

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 0.5 \text{ atm}$$

Total Pressure =  $P_1 + P_2 + P_3$

$$1.0 = 0.5 + 0.5 \rightarrow 1.0 \text{ atm}$$

*Assume both containers are 1.0 L each*

$$P_1 V_1 = P_2 V_2$$

$$(1 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 1.0 \text{ L}$$

$$P_2 = 0.5 \text{ atm} \rightarrow 0.5 \text{ atm}$$

\*Then make another volume and make are directly proportional!

8. A sealed flask contains a gas with pressure  $P$ . If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of  $P$ ?

*Doubling the moles doubles the pressure*

$$P_1 V_1 = P_2 V_2$$

*Quadrupling the volume cuts the pressure in 1/4*

$$P_1 V_1 = P_2 V_2$$

*results in an overall change in half  $P/2$*

9. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change within the container?

$$P_1 V_1 = P_2 V_2$$

*Going from 1 mole to 5 moles results in an increase of pressure by 4 times!*

$$P_1 V_1 = P_2 V_2$$

10. If a balloon containing 1000 L of gas at  $40^\circ\text{C}$  and 760 mmHg rises to an altitude where the pressure is 500 mmHg and the temperature is  $10^\circ\text{C}$ , what is the new volume of the balloon (in L)?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \rightarrow V_2 = ?$$

$$(760 \text{ mmHg} \times 1000 \text{ L}) / 313 \text{ K} = (500 \text{ mmHg} \times V_2) / 273 \text{ K}$$

$$V_2 = 501.76 \text{ L} \rightarrow 500 \text{ L}$$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$$P_1 V_1 = P_2 V_2$$

$$(87.5 \text{ mmHg} \times 500 \text{ mL}) = P_2 \times 225 \text{ mL}$$

$$P_2 = 194.44 \text{ mmHg} \rightarrow 194 \text{ mmHg}$$

12. You have 4.5 L of  $\text{NO}_2$  at a temperature of  $23^\circ\text{C}$ . What volume would the  $\text{NO}_2$  occupy at  $47^\circ\text{C}$ ? (Assume the pressure is constant).

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \rightarrow V_2 = ?$$

$$(4.5 \text{ L} \times 298 \text{ K}) / 273 \text{ K} = 4.97 \text{ L} \rightarrow 5.0 \text{ L}$$

$$V_2 = 4.97 \text{ L} \rightarrow 5.0 \text{ L}$$

13. You have a sample of a gas in a flask with a volume of 240 mL. At  $35^\circ\text{C}$  the pressure of the gas is 580 mmHg. If you decrease the temperature to  $-15^\circ\text{C}$ , what is the gas pressure at the lower temperature? *Since volume doesn't change, you can leave it out of the equation.*

$$P_1 / T_1 = P_2 / T_2$$

$$(580 \text{ mmHg} \times 273 \text{ K}) / 308 \text{ K} = P_2 \times 273 \text{ K}$$

$$P_2 = 500.48 \text{ mmHg} \rightarrow 500 \text{ mmHg}$$

# Gas Laws Homework Answer Key

**L Towne**



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