

Name \_\_\_\_\_

## Gas Laws

21. Q: What is the relationship between volume and moles in Avogadro's Principle?

A: Volume is directly proportional to the number of moles.

22. Q: Calculate the pressure of 1.0 mole of gas in a 10.0 L container at 273 K using the ideal gas law.

A: The pressure is 2.24 atm.

23. Q: According to the ideal gas law, what happens to the pressure if the volume is reduced by half while the temperature and number of moles remain constant?

A: The pressure doubles.

24. Q: If the volume of a gas is tripled and the temperature is doubled, what happens to the pressure according to the combined gas law?

A: The pressure decreases to 2/3 of its original value.

25. Q: Using gas stoichiometry, how many liters of  $\text{CO}_2$  are produced from 5.0 L of  $\text{O}_2$  in the reaction  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  at STP?

A: 5.0 L of  $\text{CO}_2$  are produced.

26. Q: In the reaction  $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ , how many liters of  $\text{NO}_2$  are produced from 4.0 L of NO at STP?

A: 4.0 L of  $\text{NO}_2$  are produced.

27. Q: Calculate the volume of  $\text{N}_2$  gas required to produce 20.0 L of  $\text{NH}_3$  in the reaction  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  at STP.

A: 10.0 L of  $\text{N}_2$  are required.

28. Q: If the temperature of a gas is increased from 300 K to 600 K and the volume is kept constant, what happens to the pressure?

A: The pressure doubles.

29. Q: Using the ideal gas law, how do you calculate the volume of gas if you know the pressure, temperature, and number of moles?

A:  $V = nRT/P$

30. Q: In the reaction  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$ , how many liters of HCl are produced from 3.0 L of  $\text{H}_2$  at STP?

A: 6.0 L of HCl are produced.



# Gas Law Study Guide Answers

**Leo J. Malone, Theodore O. Dolter**



## **Gas Law Study Guide Answers:**

## **Gas Law Study Guide Answers** Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Gas Law Study Guide Answers**," published by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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