

100. Determine the number of molecules of SO_2 molecules. Show SO_2 molecule is a trigonal planar. How many atoms of sulfur are there in 1.00 mole of SO_2 molecules?

$$\text{no. of molecules} = n \times N_A = 1.00 \times (6.02 \times 10^{23}) = 6.02 \times 10^{23} \text{ molecules}$$

$$n = \frac{m}{M} = \frac{64}{64} = 1$$

$$\text{atoms of } \text{SO}_2 = 3 \times 6.02 \times 10^{23} = 1.806 \times 10^{24} \text{ atoms}$$

101. The only commercially known diatomic sulfur are S_8 and S_2 .

(a) Describe how you can identify S_8 and S_2 .

S_8 is a cyclic molecule, whereas S_2 is a linear molecule. S_8 is a solid at room temperature, whereas S_2 is a gas. S_8 is a yellow solid, whereas S_2 is a yellow gas. S_8 is a cyclic molecule, whereas S_2 is a linear molecule. S_8 is a solid at room temperature, whereas S_2 is a gas. S_8 is a yellow solid, whereas S_2 is a yellow gas.

(b) Suggest other ways you can identify S_8 and S_2 .

S_8 is a cyclic molecule, whereas S_2 is a linear molecule. S_8 is a solid at room temperature, whereas S_2 is a gas. S_8 is a yellow solid, whereas S_2 is a yellow gas.

102. Molar conductivity is a component of molar conductivity. Molar conductivity is a component of molar conductivity.

Calculate the molar conductivity of Na_2SO_4 at 25°C.

Given: The molar conductivity of Na^+ is $50 \text{ S cm}^2 \text{ mol}^{-1}$.

$$\Lambda_m^\circ(\text{Na}_2\text{SO}_4) = 2 \times \Lambda_m^\circ(\text{Na}^+) + \Lambda_m^\circ(\text{SO}_4^{2-})$$

$$= 2 \times [50] + 80 = 180$$



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Ian Pickup



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