

How many moles of gas are contained in 22.41L at 99.85 kPa and 0 degrees C?

$$V = 22.41 \text{ L}$$

$$T = 0^{\circ}\text{C} + 273 = 273 \text{ K}$$

make sure  
your units  
match!

$$P = 99.85 \text{ kPa}$$

$$n = ?$$

$$PV = nRT$$

$$(99.85 \text{ kPa})(22.41 \text{ L}) = n \left( \frac{8.314 \text{ L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}} \right) \left( \frac{273 \text{ K}}{1} \right)$$

$$0.986 \text{ mol}$$

$$\frac{2237.6 \text{ kPa} \cdot \text{L}}{2269.7 \frac{\text{L} \cdot \text{kPa}}{\text{mol}}} = n \left( \frac{2269.7 \text{ L} \cdot \text{kPa} \cdot \text{K}}{\text{mol} \cdot \text{K}} \right)$$



# Ideal Gas Law Problems Lincoln Sudbury Regional High School

**Michael Seilmaier**





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