

# **Chemistry – Unit 5 Worksheet 2** **Empirical and Molecular Formulas**

Show all your work, show all steps, use following guidelines. Do not forget to show units and label your answers.

1. Find the empirical formula of a compound containing 20.0 g of bromine and 4.0 g of magnesium.

$$\text{magn} = \frac{4.000 \text{ g}}{24.305 \text{ g/mol}} = 0.1646 \text{ moles Mg}$$

1:1 ratio

$$\text{Br}_2 = \frac{20.000 \text{ g}}{159.808 \text{ g/mol}} = 0.1252 \text{ moles Br}_2$$



2. What is the empirical formula of a hydrocarbon compound if 1.00 g of this compound contains 0.75 g of carbon and 0.25 g of hydrogen?

$$\text{C} = 0.75 \text{ g} / 12.01 \text{ g/mol} = 0.0625 \text{ mol}$$

$$\text{hydrog} = \frac{0.250 \text{ g}}{1.008 \text{ g/mol}} = 0.248 \text{ moles H}$$

1:4 ratio

$$\text{CH}_4$$



3. A compound is analyzed and determined to contain 40.0 g of nitrogen, 40.0 g of hydrogen, and 160.0 g of oxygen. What is the empirical formula of this compound?

$$\text{N} = 40.0 \text{ g} / 14.007 \text{ g/mol} = 2.855 \text{ moles N}$$

$$\text{hydrog} = \frac{40.0 \text{ g}}{1.008 \text{ g/mol}} = 39.68 \text{ moles H}$$

1:14:2 ratio

$$\text{N}_2\text{H}_{14}\text{O}_2$$



4. A compound composed of hydrogen and oxygen is found to contain 10.0 g of hydrogen and 80.0 g of oxygen. What is the ratio of these components if it is given. Find the empirical formula.

$$\text{H} = 10.0 \text{ g} / 1.008 \text{ g/mol} = 9.92 \text{ moles H}$$

$$\text{O} = 80.0 \text{ g} / 15.999 \text{ g/mol} = 5.00 \text{ moles O}$$

10:5 ratio hydrogen to oxygen 2:1

simplest form, divide with 15.999

hydrogen = 10.000 moles hydrogen

oxygen = 5.00 moles oxygen

hydrogen = 2 moles H, oxygen = 1 mole O

# Modeling Chemistry Unit 5 Review Answers

**Jin-Ying Zhang**



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