

CP Chemistry Unit 1 Test Review KEY

Problems and questions:

- Round each of the following numbers to 3 significant figures:
a) .09882 **.0988** b) 123,459 **123,000** c) 102.45 **102** d) 2.301×10^{-2} **2.30×10^{-2}**
- Calculate, use sig figs: $(12.22 \times 14.02 \times 120.1 \times .555) / 10.00 =$ **1140**
- A sample of metal having a mass of 12.122 g is added to a graduated cylinder containing 35.2 mL of water. The water level rises to 38.4 mL. Calculate the density, use correct significant figures.
$$d = \frac{12.122}{3.2} = 3.8 \text{ g/mL}$$
- Calculate the % error for an experiment where the student measured 3.45 g and the actual value was 3.00 g?
$$\% \text{ error} = \frac{(3.45 \text{ g} - 3.00 \text{ g})}{(3.00 \text{ g})} \times 100 = 15.0\%$$
- Write the symbol for the following ions: Mg losing 2 e⁻, Al losing 3 e⁻, N gaining 3 e⁻, S gaining 2 e⁻
Mg²⁺, Al³⁺, N³⁻, S²⁻
- a) 2 b) 3 c) 12 d) 5
- List three points of Dalton's atomic theory.
A. Elements made of tiny particles, called atoms.
B. Atoms of a given element are identical; the atoms of different elements are different.
C. In chemical reactions, atoms are combined, separated, or rearranged. Atoms cannot be subdivided, created, or destroyed.
D. Atoms combine with atoms, in small whole numbers, to form compounds.
- Describe Thompson's experiment with the cathode ray tube, what did he discover?
Cathode ray tube consists of a cathode (negative) and anode (positive) side. The "ray" could be bent with magnetic or electric fields. The "ray" bent toward a positive EM field and away from a negative EM field. Discovered electron and charge to mass ratio of electron.
- Describe Millikan's Oil Drop experiment. What exactly did he determine?
He suspended a micro oil droplet in an electric field and was able to determine the actual charge of an electron to be $-1.6 \times 10^{-19} \text{ C}$. Also, determined the mass of an electron.
- Explain Rutherford's Gold Foil experiment. What were three observations and three conclusions?
He fired alpha particles at gold foil and observed: 1) most went straight through, 2) small number showed a slight deflection, 3) very few bounced straight back. He concluded: 1) mass of atom is in the nucleus, 2) nucleus is positively charged, 3) atom mostly empty space.
- What is an isotope? **Atoms of the same element that differ by # of electrons.**
- Write the symbol of the most abundant copper isotope if the average atomic mass is 63.55? **$^{64}_{29}\text{Cu}$**
- Write the nuclear symbol for an isotope with Z=40, and 45 neutrons? **$^{85}_{40}\text{Zr}$**
- Consider the diagram and match the letters with the appropriate particle
A= electron, B= neutron, C= proton
- Place the forms of hydrogen (H, H⁺, H⁻, H₂) in the table.



Atom	Cation	Anion	Molecule
H	H⁺	H⁻	H₂

Modeling Chemistry Unit 1 Review Answers

BM King



Modeling Chemistry Unit 1 Review Answers:

This book delves into Modeling Chemistry Unit 1 Review Answers. Modeling Chemistry Unit 1 Review Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry Unit 1 Review Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry Unit 1 Review Answers
- Chapter 2: Essential Elements of Modeling Chemistry Unit 1 Review Answers
- Chapter 3: Modeling Chemistry Unit 1 Review Answers in Everyday Life
- Chapter 4: Modeling Chemistry Unit 1 Review Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Modeling Chemistry Unit 1 Review Answers. The first chapter will explore what Modeling Chemistry Unit 1 Review Answers is, why Modeling Chemistry Unit 1 Review Answers is vital, and how to effectively learn about Modeling Chemistry Unit 1 Review Answers.
3. In chapter 2, the author will delve into the foundational concepts of Modeling Chemistry Unit 1 Review Answers. The second chapter will elucidate the essential principles that must be understood to grasp Modeling Chemistry Unit 1 Review Answers in its entirety.
4. In chapter 3, the author will examine the practical applications of Modeling Chemistry Unit 1 Review Answers in daily life. This chapter will showcase real-world examples of how Modeling Chemistry Unit 1 Review Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Modeling Chemistry Unit 1 Review Answers in specific contexts. The fourth chapter will explore how Modeling Chemistry Unit 1 Review Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Modeling Chemistry Unit 1 Review Answers. The final chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry Unit 1 Review Answers.

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