

## Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as  $\text{NaCl}$ . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

### Model 1 – Ion Charges for Selected Elements

1	H <sup>+</sup>											
2	Li <sup>+</sup>	Be <sup>2+</sup>						N <sup>3-</sup>	O <sup>2-</sup>	F <sup>-</sup>		
3	Na <sup>+</sup>	Mg <sup>2+</sup>	Transition elements				Al <sup>3+</sup>		P <sup>3-</sup>	S <sup>2-</sup>	Cl <sup>-</sup>	
4	K <sup>+</sup>	Ca <sup>2+</sup>	Fe <sup>2+</sup> Fe <sup>3+</sup>	Ni <sup>2+</sup> Ni <sup>3+</sup>	Cu <sup>+</sup> Cu <sup>2+</sup>	Zn <sup>2+</sup>						Br <sup>-</sup>
5	Rb <sup>+</sup>	Sr <sup>2+</sup>			Ag <sup>1+</sup>			Sn <sup>2+</sup> Sn <sup>4+</sup>				I <sup>-</sup>
6		Ba <sup>2+</sup>				Hg <sub>2</sub> <sup>2+</sup> Hg <sup>2+</sup>		Pb <sup>2+</sup> Pb <sup>4+</sup>				
	Cations							Anions				

1. Based on the information in Model 1:

- a. Identify three elements that form only one cation.  $\leftarrow$  Positive Ion  
 $\text{Na}^+$ ,  $\text{Sr}^{2+}$ ,  $\text{Al}^{3+}$  (Any box with one element works.)
- b. Identify three elements that form only one anion.  $\leftarrow$  Negative Ion  
 $\text{P}^{3-}$ ,  $\text{O}^{2-}$ ,  $\text{Br}^-$  (Any box with one element works.)
- c. Identify three elements that form more than one cation.  
 $\text{Fe}^{+2}$  and  $\text{Fe}^{+3}$ ,  $\text{Ni}^{+2}$  and  $\text{Ni}^{+3}$ ,  $\text{Cu}^{+1}$  and  $\text{Cu}^{+2}$
- d. In what region of the periodic table are these "multiple ion" elements usually located?  
Transition elements or "d" block

2. Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
- $2K^+$  and  $S^{2-}$  or  $K_2S$
3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



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**Jiben Roy**



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
























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