

Thus, the oxidation number of sodium in the compound is  $1+$  because one electron was transferred from the sodium atom. Because an electron is transferred to the chlorine atom, its oxidation number is  $1-$ .


- 7- In the chemical formula for any ionic compound, the symbol of the cation is always written first, followed by the symbol of the anion.
- 8- Subscripts, which are small numbers to the lower right of a symbol, represent the number of ions of each element in an ionic compound. If no subscript is written, it is assumed to be one.

**Practice Problems** **Chlorine** **Write Formulas**

Write formulas for the ionic compounds formed by the following ions.

19. potassium and iodide	21. aluminum and bromide
20. magnesium and chloride	22. sodium and nitride

23. **Challenge:** Write the general formula for the ionic compound formed by elements from the two groups shown on the periodic table at the right.



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**Formulas for polyatomic ionic compounds:**

- 9- polyatomic ions are ions made up of more than one atom. A polyatomic ion acts as an individual ion. Thus, the formula for a polyatomic compound follows the same rules used for a binary compound. Because a polyatomic ion exists as a unit, never change subscripts of the atoms within the ion.

# Ionic Compounds Study Guide

**James E. Brady, Fred Senese**



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## **Table of Contents Ionic Compounds Study Guide**

1. Understanding the eBook Ionic Compounds Study Guide
  - The Rise of Digital Reading Ionic Compounds Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Ionic Compounds Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ionic Compounds Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ionic Compounds Study Guide
  - Personalized Recommendations
  - Ionic Compounds Study Guide User Reviews and Ratings
  - Ionic Compounds Study Guide and Bestseller Lists
5. Accessing Ionic Compounds Study Guide Free and Paid eBooks
  - Ionic Compounds Study Guide Public Domain eBooks
  - Ionic Compounds Study Guide eBook Subscription Services
  - Ionic Compounds Study Guide Budget-Friendly Options

6. Navigating Ionic Compounds Study Guide eBook Formats
  - ePub, PDF, MOBI, and More
  - Ionic Compounds Study Guide Compatibility with Devices
  - Ionic Compounds Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ionic Compounds Study Guide
  - Highlighting and Note-Taking Ionic Compounds Study Guide
  - Interactive Elements Ionic Compounds Study Guide
8. Staying Engaged with Ionic Compounds Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ionic Compounds Study Guide
9. Balancing eBooks and Physical Books Ionic Compounds Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ionic Compounds Study Guide
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ionic Compounds Study Guide
  - Setting Reading Goals Ionic Compounds Study Guide
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ionic Compounds Study Guide
  - Fact-Checking eBook Content of Ionic Compounds Study Guide
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements



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