



# CHEMICAL BONDS



## IONIC, COVALENT, & METALLIC

**FORMS OF CHEMICAL BONDS**  
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

**Types of Chemical Bonds:**

- 1. Ionic Bond:** Formed by the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.
- 2. Covalent Bond:** Formed by the sharing of electron pairs between two atoms.
- 3. Metallic Bond:** Formed by the attraction between metal cations and a sea of delocalized electrons.

**COVALENT BONDING AND MOLECULAR STRUCTURE**

Covalent bonding occurs when two non-metal atoms share their outer electrons to form a stable molecule. The resulting molecule has a specific shape and properties.

**Diagram:** A diagram showing the formation of a covalent bond between two hydrogen atoms. Each hydrogen atom has one electron. When they combine, they share their electrons, forming a stable molecule of hydrogen gas (H<sub>2</sub>).

**Periodic Table:** A periodic table of elements is shown, with a color-coded legend for different groups of elements.

**SCIENCE REAL**

**IONIC BONDING AND MOLECULAR STRUCTURE**

Ionic bonding occurs when one atom transfers an electron to another atom, resulting in oppositely charged ions. These ions are attracted to each other, forming a stable ionic compound.

**Diagram:** A diagram showing the formation of an ionic bond between a sodium atom (Na) and a chlorine atom (Cl). The sodium atom has one valence electron, and the chlorine atom has seven valence electrons. When they combine, the sodium atom transfers its electron to the chlorine atom, forming a sodium cation (Na<sup>+</sup>) and a chloride anion (Cl<sup>-</sup>).

**Diagram:** A diagram showing the formation of a covalent bond between two oxygen atoms. Each oxygen atom has six valence electrons. When they combine, they share their electrons, forming a stable molecule of oxygen gas (O<sub>2</sub>).

**Diagram:** A diagram showing the formation of a metallic bond between metal atoms. The metal atoms are represented as cations in a sea of delocalized electrons.

**Diagram:** A diagram showing the formation of a triple covalent bond between two carbon atoms. Each carbon atom has four valence electrons. When they combine, they share three pairs of electrons, forming a stable molecule of ethyne (C<sub>2</sub>H<sub>2</sub>).

## GUIDED READING ACTIVITY

# Ionic Bonds Guided Reading And Study Answers

**Herbert Drapkin**



## **Ionic Bonds Guided Reading And Study Answers:**

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