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### Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	$Li^+$
F	7 ve	accepts $1e^-$ from Li	$F^-$

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $Li + Cl \Rightarrow LiCl$ $Li + \cdot\ddot{Cl}\cdot \rightarrow [Li]^+ [\ddot{Cl}]^-$	Li	1 ve	donates $1e^-$	$Li^+$
	Cl	1 ve	accepts $1e^-$	$Cl^-$
2) $Ca + O \Rightarrow CaO$ $Ca + \cdot\ddot{O}\cdot \rightarrow [Ca]^{+2} [\ddot{O}]^{2-}$	Ca	2 ve	donates $2e^-$	$Ca^{+2}$
	O	6 ve	accepts $2e^-$	$O^{2-}$
3) $Be + F \Rightarrow BeF_2$ $Be + 2\cdot\ddot{F}\cdot \rightarrow [Be]^{+2} [2\ddot{F}]^{2-}$	Be	2 ve	donates $2e^-$ to each F	$Be^{+2}$
	F	7 ve	each F accepts $1e^-$ each	$F^-$
4) $Mg + S \Rightarrow MgS$ $Mg + \cdot\ddot{S}\cdot \rightarrow [Mg]^{+2} [\ddot{S}]^{2-}$	Mg	2 ve	donates $2e^-$ to S	$Mg^{+2}$
	S	6 ve	accepts $2e^-$	$S^{2-}$
5) $K + F \Rightarrow KF$ $K + \cdot\ddot{F}\cdot \rightarrow [K]^+ [\ddot{F}]^-$	K	1 ve	donates $1e^-$	$K^+$
	F	7 ve	accepts $1e^-$	$F^-$

# Ionic Bonding Worksheet 1 Answer Key

**Sir Tarun Rupani**



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