

Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

WINE

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl . In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

1	H ⁺									
2	Li ⁺	Be ²⁺						N ³⁻	O ²⁻	F ⁻
3	Na ⁺	Mg ²⁺	Transition elements				Al ³⁺	P ³⁻	S ²⁻	Cl ⁻
4	K ⁺	Ca ²⁺	Fe ²⁺ Fe ³⁺	Ni ²⁺ Ni ³⁺	Cu ⁺ Cu ²⁺	Zn ²⁺				Br ⁻
5	Rb ⁺	Sr ²⁺			Ag ⁺			Sn ²⁺ Sn ⁴⁺		I ⁻
6		Ba ²⁺				Hg ²⁺ Hg ²⁺		Pb ²⁺ Pb ⁴⁺		
	Cations						Anions			

- Based on the information in Model 1:
 - Identify three elements that form only one cation.
 K^+ , Li^+ , H^+
 - Identify three elements that form only one anion.
 N^{3-} , P^{3-} , S^{2-}
 - Identify three elements that form more than one cation.
 Fe^{2+} , Ni^{2+} , Sr^{2+}
 Pb^{2+} , Ni^{3+} , Sr^{4+}
 - In what region of the periodic table are these "multiple ion" elements usually located?
 Transition elements
- Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.
 K_2S
- Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.



Ionic Bonding Pogil

Stephen W. Adams



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Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 th th The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le M ridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40

countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings

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Understanding Chemistry: Chemical bonding, 1967 *Chemical Bonding* M. S. Sethi, M. Satake, 1992 **Introductory Chemistry** Michael P. Garoutte, Ashley B. Mahoney, 2015-08-10 The ChemActivities found in Introductory Chemistry A Guided Inquiry use the classroom guided inquiry approach and provide an excellent accompaniment to any one semester Introductory text Designed to support Process Oriented Guided Inquiry Learning POGIL these materials provide a variety of ways to promote a student focused active classroom that range from cooperative learning to active student participation in a more traditional setting *Charge Ahead*, 2025 Although materials science has a long history polymers are a relatively recent addition to the world of materials Early developments focused on processing or modifying naturally occurring resins and proteins to create what we now recognize as historic plastics However demand for more accessible materials drove the development of modern plastics Over the past 150 years plastics have transformed industries and everyday life Yet today plastics often end up in undesirable places especially natural and marine environments While many are recyclable a fully circular system remains out of reach Thermosetting plastics are particularly challenging due to their irreversible chemical cross links Replacing these with reversible interactions could allow for true recyclability across all polymers This thesis explores ionic interactions modified by hydrophobic and steric screening groups as reversible cross linkers We introduce compleximers a new class of polymers that merge the recyclability of thermoplastics with the structural stability of thermosets Part I presents experimental studies on compleximers focusing on synthesis processing properties and behavior

under heat and plasticization In Chapter 2 we create compleximers by forming polyelectrolyte complexes from hydrophobically modified polyelectrolytes Once dried and desalted these materials can be hot pressed into plastics without added solvents and can be reshaped multiple times without loss of mechanical strength We also study ionic liquids as plasticizers showing they introduce a glass transition temperature while preserving solvent resistance and network integrity In Chapter 3 we examine the effects of ionic liquid plasticizers on relaxation dynamics using rheological time temperature superposition TTS fluorescence recovery after photobleaching FRAP fluorescence lifetime imaging FLIM and broadband dielectric spectroscopy BDS We find that plasticizers speed up relaxations but do not significantly alter activation energies Each method probes distinct processes from bulk relaxation rheology to small molecule mobility FRAP BDS Chapter 4 introduces a more flexible fluorine free compleximer and examines its temperature dependent modulus in the context of strong glass formers We reveal a novel inverse relationship between material fragility and relaxation spectrum width contrary to trends in literature A new scaling description connects fragility glass transition steepness and relaxation spectrum width across various ionic materials Because experimental variation of structural parameters is limited Part II uses molecular dynamics simulations to study compleximer architecture Chapter 5 explores the effects of neutral side chain length plasticizer content and charge density Reducing charge density or adding neutral chains lowers density and coordination accelerating dynamics and reducing the glass transition temperature We also compare the dynamics of ionic liquids polymerized ionic liquids and compleximers Finally Chapter 6 discusses open scientific challenges potential applications and future directions for compleximers This work highlights a promising route toward recyclable polymeric materials with robust yet reversible network structures

The Chemical Bond in Inorganic Chemistry Ian David Brown, 2002

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