

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show us the products and how to find them. The next 4 problems cover three types.

1. Find the mass of product of reaction of potassium hydroxide, KOH , and sulfuric acid, H_2SO_4 , if 100 g of potassium hydroxide, KOH , and 100 g of sulfuric acid are used.

Equation:	$\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{H}_2\text{O}$
Balance:	$2\text{KOH} + \text{H}_2\text{SO}_4 \rightarrow \text{K}_2\text{SO}_4 + 2\text{H}_2\text{O}$
Masses:	100 100 100
Mass:	50 50 100

100 g of KOH and 100 g of H_2SO_4 produce 100 g of K_2SO_4 and 100 g of H_2O .

2. How much a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. If 100 g of magnesium metal reacts with 100 g of hydrochloric acid, how much of each product is formed?

Equation:	$\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
Balance:	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
Masses:	100 100 100
Mass:	50 50 100

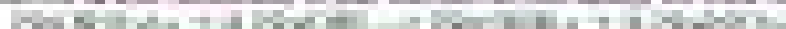
100 g of Mg and 100 g of HCl produce 100 g of MgCl_2 and 100 g of H_2 .

3. Sulfuric acid, H_2SO_4 , reacts with sodium hydroxide, NaOH , to produce sodium sulfate, Na_2SO_4 , and water. If 100 g of sulfuric acid reacts with 100 g of sodium hydroxide, how much of each product is formed?

Equation:	$\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$
Balance:	$\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
Masses:	100 100 100
Mass:	50 50 100

100 g of H_2SO_4 and 100 g of NaOH produce 100 g of Na_2SO_4 and 100 g of H_2O .

4. Chlorine gas, Cl_2 , reacts with sodium bromide, NaBr , to produce sodium chloride, NaCl , and bromine, Br_2 . If 100 g of chlorine gas reacts with 100 g of sodium bromide, how much of each product is formed?



100 g of Cl_2 and 100 g of NaBr produce 100 g of NaCl and 100 g of Br_2 .

Equation:	$\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
Balance:	$\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
Masses:	100 100 100
Mass:	50 50 100

100 g of Cl_2 and 100 g of NaBr produce 100 g of NaCl and 100 g of Br_2 .

Modeling Chemistry U8 Ws 4 V2 Answers

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