

## Chemistry – Unit 6 Notes 1 - Thomson Model of the Atom

J. J. Thomson performed experiments with cathode rays in an attempt to understand electricity – which was still a mystery in the late 1800s. Use the website *A Look Inside the Atom*<sup>1</sup> to find the conclusions that Thomson and other physicists drew regarding the mysterious cathode rays.

Thomson's 1897 Experiments - state the conclusions Thomson drew from each of his famous cathode ray experiments:

1. **First Experiment:** Thomson directed the beam at an electrometer and tried to separate the evidence of charge from the path of the beam. *What connection did Thomson find between charge and the cathode rays? Was the charge positive or negative?*
2. **Second Experiment:** Thomson tried passing the cathode ray through an electric field. *How did cathode ray beam behave when it passed through an electric field? What did he conclude after his second experiment?*
3. **Third Experiment:** Thomson did some careful measurements on how much the path of the cathode ray was bent in a magnetic field and how much energy they carried. From this work Thomson could describe the mass/charge ratio of the cathode ray particles. *What amazing result did Thomson find?*

**Thomson's Atomic Model:** Thomson presented three hypotheses from his experiments. Only two were accepted by physicists – in fact the third was shown to be wrong! From the first two came a model of the atom known as the *Plum Pudding* model. Complete the atom drawing below to illustrate Thomson's plum pudding model. Explain how this fits with his observations.



<sup>1</sup> <https://www.nip.org/history/exhibits/electron/jrays.htm>  
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# Modeling Chemistry U6 Wsv2 Answers

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## **Table of Contents Modeling Chemistry U6 Wsv2 Answers**

1. Understanding the eBook Modeling Chemistry U6 Wsv2 Answers
  - The Rise of Digital Reading Modeling Chemistry U6 Wsv2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U6 Wsv2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U6 Wsv2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U6 Wsv2 Answers
  - Personalized Recommendations
  - Modeling Chemistry U6 Wsv2 Answers User Reviews and Ratings
  - Modeling Chemistry U6 Wsv2 Answers and Bestseller Lists
5. Accessing Modeling Chemistry U6 Wsv2 Answers Free and Paid eBooks
  - Modeling Chemistry U6 Wsv2 Answers Public Domain eBooks
  - Modeling Chemistry U6 Wsv2 Answers eBook Subscription Services
  - Modeling Chemistry U6 Wsv2 Answers Budget-Friendly Options

6. Navigating Modeling Chemistry U6 Wsv2 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Modeling Chemistry U6 Wsv2 Answers Compatibility with Devices
  - Modeling Chemistry U6 Wsv2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry U6 Wsv2 Answers
  - Highlighting and Note-Taking Modeling Chemistry U6 Wsv2 Answers
  - Interactive Elements Modeling Chemistry U6 Wsv2 Answers
8. Staying Engaged with Modeling Chemistry U6 Wsv2 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry U6 Wsv2 Answers
9. Balancing eBooks and Physical Books Modeling Chemistry U6 Wsv2 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modeling Chemistry U6 Wsv2 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry U6 Wsv2 Answers
  - Setting Reading Goals Modeling Chemistry U6 Wsv2 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry U6 Wsv2 Answers
  - Fact-Checking eBook Content of Modeling Chemistry U6 Wsv2 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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