

Thermodynamics Review

1. What is a spontaneous change or reaction?

Spontaneous changes or reactions occur in a forward direction without any outside intervention.

2. Will a spontaneous reaction *always* proceed at a rapid rate towards equilibrium?

No, it will if given enough time, but thermodynamics cannot predict how fast the reaction will proceed. Kinetics is the study of reaction rates.

3. Because a reaction is spontaneous, does it mean that a lot of product will be produced when the system reaches equilibrium?

The spontaneity of a reaction does not tell you whether a lot of product will be produced or not. (A reaction that has a large negative value of ΔG , however will usually be product favored.)

We will talk more about equilibrium in Chapter 15.

4. A chemical reaction in which most of the reactants can eventually be converted to products, *if given sufficient time*, is said to be product favored.

5. If a reaction is reactant favored, does it mean that it will not occur at all?

No, it just means that not a lot of product is formed.

Chemists spend a lot of time and money varying reaction conditions to maximize the amount of product formed.

We will talk about this more next unit.

6. The thermodynamic function called entropy is a measure of the matter dispersal or disorder of the system.

7. All spontaneous processes increase the total entropy of the universe (2nd Law)

8. If the entropy of each element is taken as zero at 0 K, every substance has a finite entropy. This statement is called the 3rd Law of Thermodynamics.

PREDICT S

9. By inspection, predict whether the entropy value will increase or decrease during the following processes and give the sign of the entropy change:

a) $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{g})$ increase (+)

b) $\text{NaCl}(\text{s}) \rightarrow \text{NaCl}(\text{aq})$ increase (+)

c) $2\text{CH}_3\text{OH}(\text{l}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{g})$ increase (+)

CALCULATE S

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