

Chemistry Unit 8 Worksheet 4

Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. You can follow the problems without an answer. We want to make your final solution.

12. Calculate the number of moles of potassium chlorate, KClO_3 , that must decompose to produce potassium chloride KCl and a total of 1.5 moles of oxygen gas.



Substance	KClO_3	KCl	O_2
Moles	1.0	0	1.5
Change	-1.0	+2.0	+1.5
Initial	0	0	0

1.0 moles KClO_3 must

13. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 0.1 g of magnesium?



Substance	Mg	HCl	MgCl_2	H_2
Change	-0.004	+0.008	+0.004	+0.004
Initial	0	0	0	0

0.008 moles HCl needed

14. Ethane, C_2H_6 , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What number of moles of oxygen gas is required to react with 0.10 moles of ethane?



Substance	C_2H_6	O_2	CO_2	H_2O
Change	-0.10	+0.35	+0.20	+0.30
Initial	0	0	0	0

0.35 moles O_2 must be used to react with 0.10 moles C_2H_6

15. Sodium metal reacts with sulfuric acid to produce sodium sulfate and hydrogen gas. How many moles of sulfuric acid are needed to react with 0.10 moles of sodium metal?



Substance	Na	H_2SO_4	Na_2SO_4	H_2
Change	-0.10	+0.05	+0.05	+0.05
Initial	0	0	0	0

0.05 moles H_2SO_4 must be used to react with 0.10 moles Na

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M Tight



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New Scientist ,1987-11 *The Metallurgist and Materials Technologist* ,1982
Arthur James Wells,1974

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