

Chemistry Chapter 10

Name _____

Stoich Worksheet 2

Stoichiometry Practice Problems 1

Do this on a separate sheet of paper. Show all work and use units. Round your answers to 3 significant digits.

- What mass of iron is needed to react completely with 64.2 g of sulfur?
 Fe + S → FeS
- How much copper (I) oxide is needed to produce 75.0 g of copper (I) chloride?
 Cu₂O + 2HCl → 2 CuCl + H₂O
- 3. How much hydrogen sulfite is produced when 178 g of sulfur dioxide reacts with water?

4. If 65.0 g of iron (II) sulfide is added to hydrogen chloride, what mass of hydrogen sulfide is produced?

5. Silver bromide is a precipitate that is formed when silver nitrate reacts with sodium bromide. What mass of the precipitate is produced if you start with 48.7 g of silver nitrate?

6. How much hydrogen nitrate is needed to react completely with 36.3 g of magnesium?

$$Mg + 2HNO_3 \rightarrow Mg(NO_3)_2 + H_2$$

7. What mass of hydrogen can be produced when 82.1 g of zinc reacts with hydrogen sulfate?

For #8 & 9, you will need to use the density of oxygen gas is 1.43g/L.

- 8. What volume of oxygen is produced when 28.9g of water is decomposed? $2H_2O \rightarrow 2H_2 + O_2$
- 9. What volume of oxygen is produced from the decomposition of 145.7g of potassium chlorate?

Mixed Stoichiometry Word Problems And Answers

Marc R. Roussel

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