

- **Thermodynamics** is the study of energy and its transformations.
 - **First Law of Thermodynamics**: Energy cannot be created or destroyed, only transformed.

$$\Delta U = Q - W$$

 where ΔU is the change in internal energy, Q is the heat added to the system, and W is the work done by the system.
 - **Second Law of Thermodynamics**: The entropy of an isolated system always increases over time.

$$\Delta S \geq 0$$

 where ΔS is the change in entropy.



- **Enthalpy** (H) is a thermodynamic property that combines internal energy (U) and the product of pressure (P) and volume (V):

$$H = U + PV$$

 - **Enthalpy change** (ΔH) for a process at constant pressure is equal to the heat added to the system:

$$\Delta H = Q_p$$

 - **Enthalpy of formation** (ΔH_f°) is the enthalpy change when one mole of a compound is formed from its elements in their standard states.



- **Free Energy** (G) is a thermodynamic potential that combines enthalpy (H) and entropy (S):

$$G = H - TS$$

 where T is the temperature.
 - **Free energy change** (ΔG) determines the spontaneity of a process at constant temperature and pressure:

$$\Delta G = \Delta H - T\Delta S$$

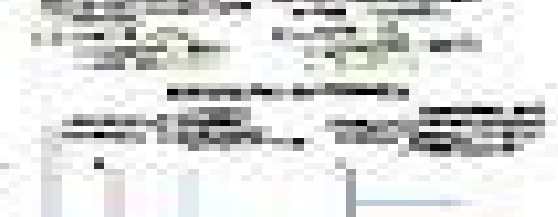
 - **Free energy of activation** ($\Delta G^\ddagger$) is the energy barrier between the reactants and the transition state.



- **Equilibrium** is a state where the forward and reverse rates of a process are equal, and the system's properties are constant.
 - **Equilibrium constant** (K) is a measure of the extent of a reaction at equilibrium.

$$K = \frac{[Products]}{[Reactants]}$$

 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change.



- **Chemical Equilibrium** is a state where the concentrations of reactants and products are constant.
 - **Equilibrium constant** (K) is a measure of the extent of a reaction at equilibrium.

$$K = \frac{[Products]}{[Reactants]}$$

 - **Le Chatelier's Principle**: If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change.



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