

Final Study Guide

- It is called the universal solvent because of its ability to dissolve most solutes
- Water molecules are held together by hydrogen bonds

Solute(s): Mixture that are homogeneous at the molecular level

Solvent: The substance being dissolved

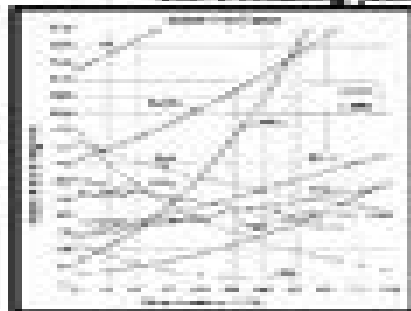
Solvent: The substance doing the dissolving, the substance there in a greater quantity

LIKE DISSOLVES LIKE: which means polar substances dissolve polar substances and nonpolar substances dissolve nonpolar substances

Notable dissolvers best...	Chemists dissolve best...
At high temperatures	At low temperatures
With stirring	At high pressures
When particles are small	With no stirring or shaking

Colligative Properties: Properties that depend on the concentration of a solution

Ex: Freezing point depression, boiling point elevation



Solubility curve: a graph used to show the dependence of solubility on temperature

unsaturated: can hold more solute, under the line

Saturated: - Won't hold any more solute, on the line

Supersaturated: - very unstable solution, crystalline ready, above the line



Electrolyte: A solution that conducts electricity
Has dissolved ions



Non-Electrolyte: A solution that does not conduct electricity
Has not dissolved ions

Acids and Bases:

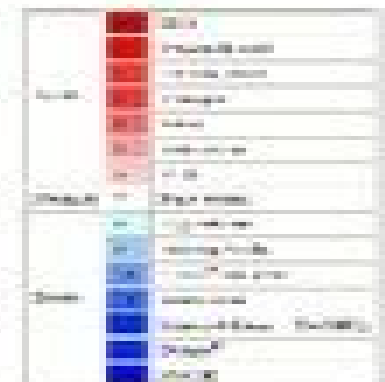
Acid: - a substance that has an excess of positive ions (H^+) have a low pH

Properties: sour, corrosive, change red litmus paper blue, neutralize when mixed with bases,

Base: - a substance that has an excess of negative ions (OH^-) have high pH

Properties: bitter, slippery, change blue litmus paper red, neutralize when mixed with acids

pH Scale: a scale that goes from 0-14, a measure of how acid or base a substance is.



Ionize: to break into positive and negative ions when dissolved in water. Things that ionize are good electrolytes



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