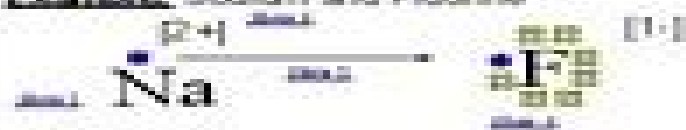


Ionic Bonding Practice 2

Instructions: Show how the following elements will ionically bond using the steps below.

1. Draw the Lewis dot structure using the number of valence electrons for each element.
Note: You might want to use different coloured pencils or dots and x's for different elements to help remember which electrons belong to which element.
2. Move the electrons and draw arrows to show the transfer of electrons to fulfill the octet rule.
3. Write the oxidation number for each ion based on how many negatively charged electrons have been given away or accepted.
4. Continue to add atoms until every atom has fulfilled the octet rule.
5. Double check and make certain all oxidation numbers add up to zero.
6. Write the chemical formula by writing the number of each element in subscript.
7. Write the name of the element. If the element has accepted electrons (anions) make the last three letters **"ide"**.

Example: Sodium and Fluorine



Octet
 $Na^+ + F^- = 0$

Chemical Formula: NaF

Chemical Name: Sodium Fluoride

1. Lithium and Oxygen



Chemical Formula: _____

Chemical Name: _____

3. Magnesium and Oxygen

Chemical Formula: _____

Chemical Name: _____

4. Lithium and Phosphorus

Chemical Formula: _____

Chemical Name: _____

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Clemens Wendtner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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