

Periodic Table: Practice Test

Modified True/False

Indicate whether the statement is true or false. If false, change the identified word or phrase to make the statement true.

- _____ 1. Protons have a negative charge. _____
- _____ 2. The modern periodic table is organized according to atomic mass. _____
- _____ 3. The horizontal rows in the periodic table are known as groups. _____
- _____ 4. The elements in a group of the periodic table have different characteristics. _____
- _____ 5. Describing a metal as ductile means that it can be pounded into a new shape. _____
- _____ 6. The most chemically reactive metals are in Group 8 of the periodic table. _____
- _____ 7. In general, the physical properties of nonmetals are different from the properties of metals. _____
- _____ 8. Atoms of the halogen family of elements typically gain two electrons when they react. _____
- _____ 9. If you know the Atomic number, you can figure out the number of protons and neutrons. _____
- _____ 10. Positively charged particles in an atom's nucleus are called neutrons. _____
- _____ 11. The valence electrons are those electrons in the outermost shell. _____
- _____ 12. Each family in the periodic table has its own characteristic properties based upon its number of valence electrons. _____
- _____ 13. The mass number is the protons plus electrons. _____
- _____ 14. Thomson's model introduced the nucleus. _____
- _____ 15. The first model to introduce energy levels was Rutherford. _____
- _____ 16. The modern day atom model is known as the Bohr model. _____
- _____ 17. Isotopes of the same element have the same number of protons. _____
- _____ 18. Isotopes of the same element have a different atomic number. _____
- _____ 19. An element has an atomic number of 12. That means it has 24 protons. _____
- _____ 20. If an element has an atomic number of 10 and a mass number of 21, the element has 10 neutrons. _____

Completion

Complete each statement.

21. The _____ is the very small center core of an atom.
22. The property of an element that indicates the number of protons in its atoms is the _____.
23. Mendeleev discovered that periodic patterns appeared when he arranged the elements in order of _____ atomic mass.

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G Psacharopoulos



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