

Fischer Esterification Project

Introduction

Esters happens to be chemical compounds that play a very significant role in the lives of people worldwide. According to *Encyclopedia.com*, naturally esters of organic acids in fruits and flowers enables them to have their distinctive odors. Mostly, they are synthesized in the laboratory and used by cosmetic industries. Nonetheless, certain evaporative esters are used as solvents for paints, lacquers and varnishes. Scientifically, common fats and oils are combination of various esters like stearin, palmitin, and linolein, formed from the alcohol glycerol and fatty acids. (*Encyclopedia.com*)

On the other hand, Fischer Esterification is a method used in synthesizing esters from carboxylic acids and alcohols. As depicted in Fig 1, it occurs with the liberation of water. The reaction is acid catalyzed and an inorganic acid as sulfuric acid can be used as a catalyst. [Williamson, 2011]. According to Le Chatelier's Principle, the reaction is reversible since the position of equilibrium can be shifted to the right by either the addition or removal of excess reactants.



Fig 1: General representation of Fischer Esterification [McMurry, 2016]

The mechanism as shown in (Fig 2) happens by the initial protonation of carboxyl group by the acid, followed by nucleophilic attack by the alcohol, proton rearrangement, loss of water and loss of catalyzing proton. The reversible reaction yields hydrolysis of the ester, [Williamson, 2011]



Fig 2: General Mechanism for Fischer Esterification [Williamson, 2011]

Objectives

For this experiment, unknown samples of a carboxylic acid and an alcohol were given for the Fischer ester synthesis. Prior to the experiment, the molar masses of the reactants to be used were given and their refractive indices were taken so all the substances were known.

Procedure

Esterification Experiment Report

M Tight



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Table of Contents Esterification Experiment Report

1. Understanding the eBook Esterification Experiment Report
 - The Rise of Digital Reading Esterification Experiment Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Esterification Experiment Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Esterification Experiment Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Esterification Experiment Report

- Personalized Recommendations
- Esterification Experiment Report User Reviews and Ratings
- Esterification Experiment Report and Bestseller Lists
- 5. Accessing Esterification Experiment Report Free and Paid eBooks
 - Esterification Experiment Report Public Domain eBooks
 - Esterification Experiment Report eBook Subscription Services
 - Esterification Experiment Report Budget-Friendly Options
- 6. Navigating Esterification Experiment Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Esterification Experiment Report Compatibility with Devices
 - Esterification Experiment Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Esterification Experiment Report
 - Highlighting and Note-Taking Esterification Experiment Report
 - Interactive Elements Esterification Experiment Report
- 8. Staying Engaged with Esterification Experiment Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Esterification Experiment Report
- 9. Balancing eBooks and Physical Books Esterification Experiment Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Esterification Experiment Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Esterification Experiment Report
 - Setting Reading Goals Esterification Experiment Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Esterification Experiment Report

- Fact-Checking eBook Content of Esterification Experiment Report
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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