

Unit 7: Balancing Chemical Reactions • Worksheet 2

Balance the following equations by inserting the proper coefficients.

1. $\underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{SO}_3$
 2. $\underline{2}\text{CH}_4 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO} + \underline{4}\text{H}_2\text{O}$
 3. $\underline{2}\text{P} + \underline{3}\text{Cl}_2 \rightarrow \underline{2}\text{PCl}_3$
 4. $\underline{2}\text{CO} + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CO}_2$
 5. $\underline{2}\text{CH}_4 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{OH}$
 6. $\underline{2}\text{Li} + \underline{1}\text{Br}_2 \rightarrow \underline{2}\text{LiBr}$
 7. $\underline{2}\text{Al}_2\text{O}_3 \rightarrow \underline{4}\text{Al} + \underline{3}\text{O}_2$
 8. $\underline{2}\text{Na} + \underline{2}\text{H}_2\text{O} \rightarrow \underline{2}\text{NaOH} + \underline{1}\text{H}_2$
 9. $\underline{6}\text{CO}_2 + \underline{6}\text{H}_2\text{O} \rightarrow \underline{1}\text{C}_6\text{H}_{12}\text{O}_6 + \underline{6}\text{O}_2$
 10. $\underline{1}\text{H}_2\text{SO}_4 + \underline{2}\text{NaCl} \rightarrow \underline{2}\text{HCl} + \underline{1}\text{Na}_2\text{SO}_4$
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
11. $\underline{3}\text{H}_2 + \underline{1}\text{SO}_2 \rightarrow \underline{1}\text{H}_2\text{S} + \underline{2}\text{H}_2\text{O}$
 12. $\underline{2}\text{CaCO}_3 + \underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CaSO}_4 + \underline{2}\text{CO}_2$
 13. $\underline{2}\text{AgNO}_3 + \underline{1}\text{CaCl}_2 \rightarrow \underline{2}\text{AgCl} + \underline{1}\text{Ca}(\text{NO}_3)_2$
 14. $\underline{2}\text{HCl} + \underline{1}\text{Ba}(\text{OH})_2 \rightarrow \underline{1}\text{BaCl}_2 + \underline{2}\text{H}_2\text{O}$
 15. $\underline{1}\text{H}_3\text{PO}_4 + \underline{3}\text{NaOH} \rightarrow \underline{1}\text{Na}_3\text{PO}_4 + \underline{3}\text{H}_2\text{O}$
 16. $\underline{1}\text{Pb}(\text{NO}_3)_2 + \underline{2}\text{KI} \rightarrow \underline{1}\text{PbI}_2 + \underline{2}\text{KNO}_3$
 17. $\underline{3}\text{CuO} + \underline{2}\text{NH}_3 \rightarrow \underline{1}\text{N}_2 + \underline{3}\text{Cu} + \underline{3}\text{H}_2\text{O}$
 18. $\underline{1}\text{C}_2\text{H}_5\text{OH} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{3}\text{H}_2\text{O}$
 19. $\underline{2}\text{C}_2\text{H}_6 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{COOH} + \underline{2}\text{H}_2\text{O}$
 20. $\underline{3}\text{NO}_2 + \underline{1}\text{H}_2\text{O} \rightarrow \underline{2}\text{HNO}_3 + \underline{1}\text{NO}$

Modeling Chemistry Worksheet Rearranging Atoms

Answer Key

**Robert C. Smoot, Smoot, Richard G.
Smith, Jack Price**



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Table of Contents Modeling Chemistry Worksheet Rearranging Atoms Answer Key

1. Understanding the eBook Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - The Rise of Digital Reading Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Worksheet Rearranging Atoms Answer Key

- Personalized Recommendations
 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key User Reviews and Ratings
 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key and Bestseller Lists
5. Accessing Modeling Chemistry Worksheet Rearranging Atoms Answer Key Free and Paid eBooks
 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key Public Domain eBooks
 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key eBook Subscription Services
 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Modeling Chemistry Worksheet Rearranging Atoms Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Highlighting and Note-Taking Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Interactive Elements Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 8. Staying Engaged with Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 9. Balancing eBooks and Physical Books Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Setting Reading Goals Modeling Chemistry Worksheet Rearranging Atoms Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Modeling Chemistry Worksheet Rearranging Atoms Answer Key

- Fact-Checking eBook Content of Modeling Chemistry Worksheet Rearranging Atoms Answer Key
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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