

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

Use an appropriate chemical equation, write a complete and balanced equation for each of the problems below. Show the procedure and final answer. The units for each case listed appear.

1. Calculate the number of moles of potassium hydroxide,  $\text{KOH}$ , and the mass of potassium hydroxide,  $\text{KOH}$ , and 1.00 moles of oxygen gas.



2 moles KOH  
 1 mole  $\text{O}_2$

2. How many grams of hydrogen gas are produced when 10.0 grams of hydrogen peroxide decompose? (Molar masses:  $\text{H}_2\text{O}_2 = 34.0 \text{ g/mol}$ ,  $\text{H}_2 = 2.0 \text{ g/mol}$ )



2.00 grams  $\text{H}_2$

3. How many grams of  $\text{Fe}_2\text{O}_3$  are produced when 10.0 grams of iron react with oxygen gas? (Molar masses:  $\text{Fe} = 55.8 \text{ g/mol}$ ,  $\text{Fe}_2\text{O}_3 = 159.7 \text{ g/mol}$ )



10.0 grams  $\text{Fe}_2\text{O}_3$

4. Calculate the number of moles of oxygen gas produced when 10.0 moles of  $\text{KClO}_3$  decompose.



15.0 moles  $\text{O}_2$

# Modeling Chemistry Unit 8 2 Answers

**RJ Alexander**



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