10 • Orbital Hybridization & Molecular Orbitals

PRACTICE TEST

- Which hybridization is associated with a steric number of 37
 - 2 Ce 98 18

- F-B-FI
- The molecule BrF₃ has a steric number of on the central atom?
 - **63** 3
- b) 4

- What is the hybridization of Br in BrF 1?

(d) sp'd SN = 5

b) sp

- e) sp
- How many equivalent sp³d orbitals are there?
 - m) 3

- do 6
- What type of hybridization is associated with a square planar molecular shape?
 - in b see !

do sp'di

by sp2

- ch sp
- What shape for electron pairs is associated with sp³d² hybridization?
 - ia) limear

- tetrahedrali
- b) square planar
- octahedral

- c) bent
- What hybridization is predicted for phosphorus in the PCI₁ molecule?
 - p) sp^2

- Letreholal superindal

- A double bond contains sigma bond(s) and (pi bond(s).
 - a) 0.2

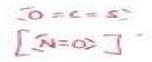
- c) 2,0 (d) 1,1
- What angle exists between orbitals in sp3d2 hybrid perhitals/2
 - 90.00 180.0*

- 109.5"
- 10. Which of the following elements is most likely to display sp³d hybridization? Su = 3
 - at oxygen
- bb- mitrogen

- c) phosphorus
- How many sigma (σ) and pi (π) electrons pairs are
 - a) four σ and zero π d) two σ and four π
 - b) three σ and two π e) one σ and three π
 - (c)) two o and two x
- 12. What is the hybridization of the oxygen atoms in
- (0-64 CH₃OH and CO₂, respectively? (0=C = 0)
- d-€-µ a) sp', sp'
- d) sp², sp² (Su = 3

- - All of the following species contain two π-bonds EXCEPT
 - a) SCN 16 a-
- b) CO IDe-
- c) H₂CCO Hear





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Thomas Griffiths

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