

$$\text{percent yield} = \frac{\text{actual yield (g)}}{\text{theoretical yield (g)}} \times 100\%$$

↳ from B-E-A chart

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation.
- Identify the given (with units) and what you want to find (with units).
- Show set up with units. Give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



1 mol  $\text{H}_2\text{O}$  : 18.02 g/mol

36 g  $\frac{1\text{ mol H}_2\text{O}}{18.02\text{ g}} = 1.99\text{ mol H}_2\text{O}$

1 mol  $\text{H}_2$  : 2.02 g

2 mol  $\text{H}_2 = \boxed{4.04\text{ g H}_2}$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



1 mol  $\text{NaCl} = 58.44\text{ g}$   
 $581\text{ g NaCl} \times \frac{1\text{ mol NaCl}}{58.44\text{ g}} = 9.94\text{ mol NaCl}$

1 mol  $\text{Na} = 22.99\text{ g}$

$9.94\text{ mol Na} \times 22.99\text{ g} = \boxed{228.52\text{ g Na}}$

# Modeling Chemistry U8 V2 Answers

**Michael Hoelscher**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Modeling Chemistry U8 V2 Answers:**

This book delves into Modeling Chemistry U8 V2 Answers. Modeling Chemistry U8 V2 Answers is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Modeling Chemistry U8 V2 Answers, encompassing both the fundamentals and more intricate discussions.

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- Chapter 2: Essential Elements of Modeling Chemistry U8 V2 Answers
- Chapter 3: Modeling Chemistry U8 V2 Answers in Everyday Life
- Chapter 4: Modeling Chemistry U8 V2 Answers in Specific Contexts
- Chapter 5: Conclusion

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