

Answer all the questions in the spaces provided.

- 1 (a) The solubility of the Group 2 sulfates decreases down the group.

Explain this trend.

.....  
 .....  
 .....  
 ..... [3]

- (b) Describe what is observed when magnesium and barium are reacted separately with an excess of dilute sulfuric acid.

magnesium .....

barium .....

[1]

- (c) The solubility product,  $K_{sp}$ , of  $\text{BaSO}_4$  is  $1.08 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$  at 298 K.

Calculate the solubility of  $\text{BaSO}_4$  in g per 100  $\text{cm}^3$  of solution.

solubility of  $\text{BaSO}_4$  = ..... g per 100  $\text{cm}^3$  of solution [2]

# Modeling Chemistry U6 Test

**M Lipman**



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