

U3 - Quest 2 b - Electron Configurations & Bohr Models			
4	3	2	1
15-16 pts	13-14	10-12	9 pts or less

1. What is the full electron configuration for Indium (In, atomic number = 49)?
2. Draw a Bohr model of an Indium-115 ( $^{115}\text{In}$ ) atom at ground state. Indicate the number of the protons, neutrons, total electrons, and the number of electrons located in each energy level. Label the nucleus, and energy levels ( $n = 1, n = 2$ , etc.).

#p+ \_\_\_\_\_

#n+ \_\_\_\_\_

#e- \_\_\_\_\_

[total]

 $n = 1$  \_\_\_\_\_ e- $n = 2$  \_\_\_\_\_ e- $n = 3$  \_\_\_\_\_ e- $n = 4$  \_\_\_\_\_ e- $n = 5$  \_\_\_\_\_ e-

3. Write the condensed (noble gas) configurations for each of the following elements. Then determine the number of valence (outermost) electrons in each element.

Element	Atomic Number	Condensed Electron Configuration	Number of Valence Electrons
Sulfur, S	16		
Scandium, Sc	21		
Cesium, Cs	55		

4. How many valence electrons are in each of the following elements:

Bi (83) \_\_\_\_\_

Ca (20) \_\_\_\_\_

Kr (36) \_\_\_\_\_

As (33) \_\_\_\_\_

# Modeling Chemistry U3 Obj V2 0 Answers

**Tobias Bleicker**



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