

Test Quantities in Chemical Reactions SCH3U #13

Name: _____

____K ____T ____C ____A

Assessment of Learning

(Product)

10

20

8

7

Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of AuCl_3 ?

A) 96 g

B) 130 g

C) 232.5 g

D) 303.6 g

2. Determine the mass in grams of 0.0488 mol cobalt.

A) 0.00852 g

B) 0.294 g

C) 2.88 g

D) 0.00830 g

3. The number of molecules of hydrogen peroxide found in 3.15 mol H_2O_2 is

A) 3

B) 1.90×10^{21}

C) 3.80×10^{21}

D) 7.60×10^{21}

4. The percentage of oxygen, by mass, in $\text{Mg}_3(\text{PO}_4)_2$ is

A) 15.9%

B) 23.8%

C) 27.7%

D) 48.7%

E) None of the above

5. An unknown gas is found to consist of 10.48 g of nitrogen and 11.96 g of oxygen. What is the percentage composition of this gas?

A) 28.6% N, 71.4 % O

B) 48.7%N, 51.3% O

C) 53.2% N, 46.7% O

D) 74.5% N, 25.5% O

6. A compound contains 38.7% K, 13.9% N, and 47.4% O by mass. What is the empirical formula of the compound?

A) KNO_3

B) $\text{K}_2\text{N}_2\text{O}_3$

C) KNO_2

D) $\text{K}_2\text{N}_2\text{O}_2$

E) K_4NO_3

7. If the molar mass of a hydrocarbon is 26.0 g/mol, and its empirical formula is CH , its molecular formula is

A) CH

B) C_2H_2

C) C_3H_3

D) C_4H_4

E) C_6H_6

8. Consider the following reaction: $\text{Ca}(\text{OH})_{2(aq)} + \text{Al}_2(\text{SO}_4)_{3(aq)} \rightarrow \text{CaSO}_{4(s)} + \text{Al}(\text{OH})_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

A) 1:1

B) 2:1

C) 3:3

D) 1:2

E) 3:1

9. The limiting reactant in a chemical reaction is the _____, that is _____.

A) Reactant, in excess

B) Product, left over

C) Reactant, used up

D) Product, used up

Modeling Chemistry Mole Test Answer

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Table of Contents Modeling Chemistry Mole Test Answer

1. Understanding the eBook Modeling Chemistry Mole Test Answer
 - The Rise of Digital Reading Modeling Chemistry Mole Test Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Mole Test Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Mole Test Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Mole Test Answer
 - Personalized Recommendations
 - Modeling Chemistry Mole Test Answer User Reviews and Ratings
 - Modeling Chemistry Mole Test Answer and Bestseller Lists
5. Accessing Modeling Chemistry Mole Test Answer Free and Paid eBooks
 - Modeling Chemistry Mole Test Answer Public Domain eBooks
 - Modeling Chemistry Mole Test Answer eBook Subscription Services
 - Modeling Chemistry Mole Test Answer Budget-Friendly Options

6. Navigating Modeling Chemistry Mole Test Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Chemistry Mole Test Answer Compatibility with Devices
 - Modeling Chemistry Mole Test Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Chemistry Mole Test Answer
 - Highlighting and Note-Taking Modeling Chemistry Mole Test Answer
 - Interactive Elements Modeling Chemistry Mole Test Answer
8. Staying Engaged with Modeling Chemistry Mole Test Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Mole Test Answer
9. Balancing eBooks and Physical Books Modeling Chemistry Mole Test Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Mole Test Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Mole Test Answer
 - Setting Reading Goals Modeling Chemistry Mole Test Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Mole Test Answer
 - Fact-Checking eBook Content of Modeling Chemistry Mole Test Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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