



## Chemistry Chapter 10

Name \_\_\_\_\_

### Stoich Worksheet 2

#### Stoichiometry Practice Problems 1

Do this on a separate sheet of paper. **Show all work and use units.** Round your answers to 3 significant digits.

1. What mass of iron is needed to react completely with 64.2 g of sulfur?



2. How much copper (I) oxide is needed to produce 75.0 g of copper (I) chloride?



3. How much hydrogen sulfite is produced when 178 g of sulfur dioxide reacts with water?



4. If 65.0 g of iron (II) sulfide is added to hydrogen chloride, what mass of hydrogen sulfide is produced?



5. Silver bromide is a precipitate that is formed when silver nitrate reacts with sodium bromide. What mass of the precipitate is produced if you start with 48.7 g of silver nitrate?



6. How much hydrogen nitrate is needed to react completely with 36.3 g of magnesium?



7. What mass of hydrogen can be produced when 82.1 g of zinc reacts with hydrogen sulfate?



For #8 & 9, you will need to use the density of oxygen gas is 1.43g/L.

8. What volume of oxygen is produced when 28.9g of water is decomposed?



9. What volume of oxygen is produced from the decomposition of 145.7g of potassium chlorate?



# Mixed Stoichiometry Word Problems And Answers

**Petrogav International Oil & Gas  
Training Center**



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