

## Unit 7 Chemistry - Nail Lab

**Purpose:** After you have completed the lab, think about what was the overall goal for performing the lab. What were we trying to accomplish?

### Procedure:

#### Step 1

1. Mass a clean & labeled \_\_\_\_\_ mL beaker.
2. Put between \_\_\_\_\_ g of copper (II) chloride in the beaker. Record the mass of the beaker and the copper (II) chloride.
3. Add \_\_\_\_\_ mL of distilled water to the beaker. Stir to dissolve the solid.
4. Obtain the mass of \_\_\_\_\_ iron nails together to ~ 0.5 g.
5. Place the nails in the copper (II) chloride solution. Observe the reactions; record your observations. Place the labeled beaker in the place designated by your teacher.

#### Step 2

- a. Remove the nails. Rinse or scrape the \_\_\_\_\_ (copper metal) from the nails into your labeled beaker. Place the nails in a separate labeled small beaker. Note the appearance of the nails in your observations. Place this beaker in the drying oven.
- b. \_\_\_\_\_ the solution from the 250 mL beaker. Rinse the precipitate with about \_\_\_\_\_ mL of distilled water. Try to lose as little of the solid copper as you can when you decant. Add a 2<sup>nd</sup> rinse with distilled water, rinse the copper with \_\_\_\_\_ . Rinse your last time with distilled water. Then place the labeled beaker in the drying oven.
- c. Mass the dry nails, and then discard the nails.

#### Step 3

- a. Mass the beaker + dry copper. Discard the copper in the place designated by your teacher. Wash your beaker and let dry.

### Data:

Mass 250 mL beaker	
Mass 250 mL beaker + copper (II) chloride	
Mass nails before reaction	
Mass nails after reaction	

# Modeling Chemistry U7 Nail Lab V2 0 Answer Key

**S Ashworth**



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