

Chemistry – Unit 6 Worksheet 1

After having observed molecular models which show 3D spatial arrangements and discussed the spatial geometry, based on, formal positively charged ions, local formal, which, more spatial geometry, based on formal negatively charged ions (anions). However, rather than a single central metal coordination spatial geometry, we substituted the formal coordination theory the real molecular effects to be and positive. After all, these examples are positive that avoid that that values for which these elements combine to make the complexes that change off their lone pair of electrons.

1. Write the formula and draw the spatial arrangement that results from the following. The name and ions are given in parentheses to guide.

3D spatial arrangement	1. octahedral 2. nitrogen	2. tetrahedral 3. nitrogen	2. octahedral 3. sulfur	2. octahedral 3. sulfur
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spatial arrangement



3D spatial arrangement	2. tetrahedral 3. nitrogen	1. trigonal bipyramidal 3. nitrogen	2. octahedral 3. sulfur
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spatial arrangement



3D spatial arrangement	2. trigonal bipyramidal 3. sulfur	2. tetrahedral 3. fluorine	2. trigonal bipyramidal 3. fluorine	2. tetrahedral 3. fluorine
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spatial arrangement



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