HA+ NOOH -> NOA+ + +100
c) [NaOH] (in mol/L), x md NaOH = 0.0500Lx 0.00690md x 1 md NaOH = 1.73x10 4 md NaOH
[NaOH] = 1.73×10 mod = 0.0081 M
d) if the resulting solution is acidic, basic or neutral at the equivalence point, and A- is weakly basic about 88 thm
e) an appropriate indicator for the titration.
Phenolphthalein
8. Triethanolamine, a weak monoprotic base, is titrated with HCl. The following data are obtained.
 20.00 mL of triethanolamine 16.30 mL of HCI Initial pH of acid = 10.355
$pH_{\frac{1}{2}}$ = 7.763 at addition of 8.15 mL of acid during titration $pOH_{\frac{1}{2}}$ $/4-7.73$ = 6.237 Determine a) the K_b for triethanolamine,
Kb= LOH-Jz
= 10-polly FIHS should read instal [trichankamine] &
b) initial (ms) (in mol/L). pothom) = 14.000 - 10355 = 3.645 Ka = [OH][HA] = 5.79x10-7
A+ HD = OH + HA 7-2.2x10-4. (2.26x10-4)2 17 597x10-7
C-226×6" +2.7810" +2286-4 X=0.0859 + 0.0002/2
X=0.0861M

Modeling Chemistry U6 Review V1 Answers

AW Rasmussen

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