## Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| Equation: | 2 KClO <sub>3</sub> → | 2 KCI + | 3 O2 |
|-----------|-----------------------|---------|------|
| Before:   | 1.2                   | •       | 0    |
| Change:   | -1.2                  | +1.2    | +1.8 |
| After:    | 0                     | 1.2     | 1.8  |

1.2 moles KClO<sub>3</sub> needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| Equation: | Mig -  | 2 HCl →  | MgCl <sub>2</sub> + | FE 2  |
|-----------|--------|----------|---------------------|-------|
| Before:   | 0.100  | 7        | 0                   | 0     |
| Change:   | -0.100 | -0.200 + | 0.160 +             | 0.100 |
| Afters    | 0      | .2       | 0.100               | 0.100 |

0.200 moles HCI needed

0.73 g x 1 mole = .0040 mole 183 g

3. Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| Equation: | 2 C2H6 | + 7 O <sub>2</sub> | → 4 CO; | + 6 H <sub>2</sub> O |
|-----------|--------|--------------------|---------|----------------------|
| Before:   | 2.20   | 26.26              | 0       | 0                    |
| Change:   | -2.20  | - 7.70             | + 4,40  | + 6.60               |
| After:    | 0      | 2.5                | 4.40    | 6.60                 |

O<sub>2</sub> needed: 7.70 mole x 32.0 g = 246 g O<sub>2</sub> required I mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:

| Ni(NO <sub>3</sub> ) <sub>2</sub> + 2 NaOH> Ni(OH) <sub>2</sub> + 2 NaNO <sub>3</sub> |                |        |                       | 0,        |  |
|---|----------------|--------|-----------------------|-----------|--|
| Equation:   | $Ni(NO_3)_2 +$ | 2 NaOH | → Ni(OH) <sub>2</sub> | + 2 NaNO3 |  |
| Before:   | 0.0040         | XS     | 0                     | 0         |  |

| The state of the s |          | The Property of the Section 1972 | The second second                        |
|--|----------|----------------------------------|--|
| 0.0040   | NS       | 0                                | 0  |
| - 0.0040   | - 0.0080 | +0.0040                          | +0.0080                                  |
| 0  | NS       | 0.0040                           | 0.0080                                   |
|  | 0.0040   | 0.0040 XS                        | 0.0040 XS 0<br>- 0.0040 - 0.0080 +0.0040 |

NaNO<sub>3</sub> produced: 0.0080 mole x <u>85.0 g</u> = 0.68 g 1 mole

# **Modeling Chemistry Unit 4 4 Answers**

**L** Reisser

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