

Stoichiometry Worksheet 2: Percent Yield

What aspects of the procedure should be:

- minimized to minimize experimental error?
- maximized to maximize the amount of product?
- minimized to minimize the amount of waste?

1. Calculate the theoretical yield of the product from the amount of reactant used. (Theoretical yield is the maximum amount of product that can be formed from a given amount of reactant.)

$$\text{Theoretical Yield} = \frac{\text{Mass of Reactant}}{\text{Molar Mass of Reactant}} \times \text{Molar Mass of Product}$$

$$\text{Theoretical Yield} = \frac{10.0 \text{ g}}{100.0 \text{ g/mol}} \times 100.0 \text{ g/mol} = 10.0 \text{ g}$$

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2. Calculate the percent yield of the product from the actual yield and the theoretical yield. (Percent yield is the ratio of the actual yield to the theoretical yield, expressed as a percentage.)

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

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3. Calculate the percent yield of the product from the actual yield and the theoretical yield. (Percent yield is the ratio of the actual yield to the theoretical yield, expressed as a percentage.)

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4. Calculate the percent yield of the product from the actual yield and the theoretical yield. (Percent yield is the ratio of the actual yield to the theoretical yield, expressed as a percentage.)

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