19.0 g L. 2003

Answer the following questions about theoretical yield and percent yield. Make sure you show all work, units, and round for sig figs.

1. Use the following equation to answer the questions below.

$$N_{2 (g)} + 3 H_{2 (g)} \rightarrow 2 NH_{3 (g)}$$

If I start this reaction with 57.0 grams of nitrogen with excess hydrogen, how many grams of ammonia (NH<sub>3</sub>) can theoretically be produced?

If 60.0 grams of ammonia is actually produced, what was my percent yield of ammonia?

Use the following equation to answer the questions below.

If 11.3 grams of carbon dioxide gas is used in this reaction, how many grams of lithium carbonate can theoretically be produced?

If 14.4 grams of lithium carbonate are actually produced, what was my percent yield?

# **Modeling Chemistry Stoichiometry 2 Percent Yield Answers**

**RC Schank** 

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