

# Using Solubility Rules

- Soluble: will readily dissolve in water
- Insoluble: will not dissolve in water
- Slightly soluble: a tiny amount will dissolve in water

Soluble Compounds	Insoluble Compounds
$\text{NO}_3^-$ , $\text{Na}^+$ , $\text{K}^+$ , $\text{NH}_4^+$	$\text{S}^{2-}$ , $\text{CO}_3^{2-}$ , $\text{PO}_4^{3-}$
$\text{Cl}^-$ , $\text{Br}^-$ , $\text{I}^-$ ( $\text{Ag}^+$ , $\text{Hg}^{2+}$ , $\text{Pb}^{2+}$ )	$\text{OH}^-$ ( $\text{Na}^+$ , $\text{K}^+$ , $\text{Ca}^{2+}$ )
$\text{SO}_4^{2-}$ ( $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , $\text{Ca}^{2+}$ )	

# Model 1 Rules Of Solubility In Aqueous Solutions Answers

**The Princeton Review**



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