

# Chapter 11: Liquids and Solids

## A. Intermolecular Forces

1. The addition of heat changes the state of a compound from solid to liquid to gas as kinetic and potential energy are added to the molecules
  - a. Kinetic Energy involves the motion of the molecules
  - b. Potential Energy involves how far apart the molecules are



2. Intermolecular Forces hold the molecules together
  - a. Ion-Dipole Interactions:  $\text{Na}^+\text{Cl}^-$  interactions with water
  - b. Dipole-Dipole Attractions: interaction of polar molecules
  - c. Hydrogen Bonding: special dipole-dipole attraction in molecules where H is bonded to an electronegative atom (O, N, F, etc...)
    - i. Small size of H lets the dipoles get very close
    - ii. Large electronegativity difference makes attraction great

# Intermolecular Forces Projects

**Julia Schneider**



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