

The Hydrogen Atom – Student Challenge

Background Material

Carefully read the background-information Energy Levels, Light, and Transitions, and answer the following questions as clearly (and completely) as you can.

Operation 1: Compare the following table which compares how the Bohr Model and the Quantum-mechanical approach the Hydrogen atom. In each “describing” box describe things as accurately as you can assuming they are true.

Concept	Bohr Model	Quantum Model
Electrons	positively charged particles in flat orbits	positively charged particles in flat orbits
Electrons	orbiting electrons	negatively charged clouds
Orbitals	Electrons do not move, they give the probability picture	Electrons, being particles without mass
Transitions of electrons to the first orbital	emitting a photon as the Bohr radius	Electron wave
Energy of the lowest energy level	$E_1 = -13.6 \text{ eV}$	It is about?

Operation 2: Compare the following table which compares the properties of two different elements by relating their approximate relationships. The data now is completed for you. The red photon has a longer wavelength, smaller frequency, smaller energy, and the same velocity. How do you see an orbiting electron?

Hydrogen Atom Student Guide Solutions Naap

K Morrison



Hydrogen Atom Student Guide Solutions Naap:

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