

- 2 -

1. Which of the following processes provide evidence of the particulate nature of matter?

- I. Capillarity
- II. Diffusion
- III. Filtration
- IV. Osmosis

- (A) I and III only
- (B) I and IV only
- (C) II and III only
- (D) II and IV only

2. 'Mass number' is the number of

- (A) neutrons plus protons
- (B) neutrons minus protons
- (C) electrons plus neutrons
- (D) electrons plus protons

3. Which of the following elements has 7 electrons in its outer shell?

- (A) Hydrogen
- (B) Oxygen
- (C) Nitrogen
- (D) Chlorine

4. An atom that has 20 electrons when it ionizes has the same electronic configuration as

- (A) argon
- (B) calcium
- (C) neon
- (D) sodium

5. $^{23}_{11}\text{Na}$ and $^{24}_{11}\text{Na}$ are isotopes because they

- (A) conduct electricity
- (B) react vigorously with water
- (C) have the same number of protons
- (D) have the same relative atomic mass

6. Sulfur and oxygen are in the same group of the periodic table because

- (A) they can react with each other
- (B) they can form covalent compounds
- (C) they have the same number of electrons in their outer shell
- (D) the atomic number of sulfur is 16 and the relative atomic mass of oxygen is 16

Item 7 refers to the following quantities of atoms.

- I. 12 g of carbon
- II. 16 g of sulfur
- III. 23 g of sodium
- IV. 12 g of magnesium

(Relative mass number: C = 12; S = 32; Mg = 24; Na = 23)

7. Which of the following pairs represents equal quantities of atoms?

- (A) I and II only
- (B) I and IV only
- (C) II and III only
- (D) II and IV only

8. An ion with a single negative charge may be converted into a neutral atom by

- (A) losing a neutron
- (B) gaining a proton
- (C) losing an electron
- (D) gaining an electron

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 - The Rise of Digital Reading Gcse Physics June 2015 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Physics June 2015 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Physics June 2015 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Physics June 2015 Answers
 - Personalized Recommendations
 - Gcse Physics June 2015 Answers User Reviews and Ratings
 - Gcse Physics June 2015 Answers and Bestseller Lists

5. Accessing Gcse Physics June 2015 Answers Free and Paid eBooks
 - Gcse Physics June 2015 Answers Public Domain eBooks
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 - Gcse Physics June 2015 Answers Budget-Friendly Options
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 - Gcse Physics June 2015 Answers Compatibility with Devices
 - Gcse Physics June 2015 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gcse Physics June 2015 Answers
 - Highlighting and Note-Taking Gcse Physics June 2015 Answers
 - Interactive Elements Gcse Physics June 2015 Answers
8. Staying Engaged with Gcse Physics June 2015 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Physics June 2015 Answers
9. Balancing eBooks and Physical Books Gcse Physics June 2015 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Physics June 2015 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Physics June 2015 Answers
 - Setting Reading Goals Gcse Physics June 2015 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Physics June 2015 Answers
 - Fact-Checking eBook Content of Gcse Physics June 2015 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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