

Unit 8 Worksheet 3: Molar Concentration

Part A: Describing Molar Concentration

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?
2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate.
3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.
4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.
5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

Holt Chemistry Concentration And Molarity Answer

G.H. Duffey



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 **Principles of Modern Chemistry** David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses **PRINCIPLES OF MODERN CHEMISTRY** continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom **Modern Chemistry**, 1990 *Modern Chemistry* H. Clark Metcalfe, Joseph F. Castka, 1966 Fundamentals of Chemistry: A Modern Introduction (1966) Frank Brescia, 2012-12-02 **Fundamentals of Chemistry A Modern Introduction** focuses on the formulas processes and methodologies used in the study of chemistry The book first looks at general and historical remarks definitions of chemical terms and the classification of matter and states of aggregation The text then discusses gases Ideal gases pressure of a gas confined by a liquid Avogadro's Law and Graham's Law are described The book also discusses aggregated states of matter atoms and molecules chemical equations and arithmetic thermochemistry and chemical periodicity The text also highlights the electronic structures of atoms Quantization of electricity spectra of elements quantization of the energy of an electron associated with nucleus the Rutherford Bohr nuclear theory hydrogen atom and representation of the shapes of atomic orbitals are explained The text also highlights the types of chemical bonds hydrocarbons and their derivatives intermolecular forces solutions and chemical equilibrium The book focuses as well on ionic solutions galvanic cells and acids and bases It also discusses the structure and basicity of hydrides and oxides The reactivity of hydrides charge of dispersal and basicity effect of anionic charge inductive effect and basicity and preparation of acids are described The book is a good source of information for readers wanting to study chemistry **Practical Environmental Analysis** Miroslav Radojevic, Vladimir N Bashkin, 2015-11-09 New techniques improved understanding and changes in regulations relating to environmental analysis means that students technicians and lecturers alike need an up to date guide to practical environmental analysis This unique book provides detailed instructions for practical experiments in environmental analysis The comprehensive coverage includes the chemical analysis of important pollutants in air water soil and plant tissue and the experiments generally require only basic laboratory equipment and instrumentation The content is supported by theoretical material explaining amongst other concepts the principles behind each method and the importance of various pollutants Also included are suggestions for projects and worked examples

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Liquids, Solutions, and Interfaces W. Ronald Fawcett, 2004-07-01 Fifty years ago solution chemistry occupied a major fraction of physical chemistry textbooks and dealt mainly with classical thermodynamics phase equilibria and non equilibrium phenomena especially those related to electrochemistry Much has happened in the intervening period with tremendous advances in theory and the development of important new experimental techniques This book brings the reader through the developments from classical macroscopic descriptions to more modern microscopic details *Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb* Wade A. Freeman, 1986

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