Name Key Class Date

Acid-Base Theories

Summarize the three main acid-base theories in the table below.

	ACID	BASE
Arrhenius	H° or H ₂ O° producer in aqueous solution	OH producer in aqueous solution
Brensted-Lowry	proton (H°) dener	proton (H*) acceptor
Lewis	electron-pair acceptor	electron-pair donor

Use the following chemical equations to answer question 2.

$$HC1 + H_2O \rightarrow H_1O^+ + C1^-$$

According to Arrhenius, HCl would be classified as a(n) acid and NH₃ would be classified as a(n) base. Explain why.

HCl produces hydronium ions (H₂O*) in water. NH₂ produces hydroxide ions (OH ²) in water.

For problems 3 - 5, label the acid (A), base (B), conjugate acid (CA), and conjugate base (CB) in each of the following reactions. Show the transfer of the proton (H') by drawing an arrow.

3.
$$H_1SO_1 + NH_2 \rightarrow HSO_1 + NH_4$$

A B CB CA

4.
$$HC_2H_3O_2 + H_2O \rightarrow H_3O' + C_2H_3O_2$$

A B CA CB

For problems 6 - 9, give the conjugate base for each Bronsted-Lowry acid.

For problems 10 - 13, give the conjugate acid for each Bronsted-Lowry base.

Holt Chemistry Acid And Bases Test Answers

IM Harris

Holt Chemistry Acid And Bases Test Answers:

Unveiling the Power of Verbal Art: An Mental Sojourn through Holt Chemistry Acid And Bases Test Answers

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