

Holt Modern Chemistry Review

CHAPTER 6: CHEMICAL BONDING

The following pages contain the bulk (but not all) of the information for the chapter 6 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test.
***** Test items will be recall, examples, and/or application of this content. *****

OUTCOMES

- Collaborate with peer(s) to understand chemistry content (C C)
- Communicate chemistry content to teacher and peer(s) (E C)
- 6.2: Use the Lewis dot notation, illustrate the covalent bond as the sharing of electrons (F & PK)
- 6.3: List important physical properties of ionic compounds; sum of oxidation numbers = zero (F & PK)

6.1: INTRODUCTION TO CHEMICAL BONDING

- **Vocabulary**
 - **chemical bond** – the attractive force that holds atoms or ions together
 - **ionic bonding** – a force that attracts electrons from one atom to another, which transforms a neutral atom into an ion
 - **covalent bonding** – a bond formed when atoms share one or more pairs of electrons
 - **nonpolar-covalent bond** – a covalent bond in which the bonding electrons are equally attracted to both bonded atoms
 - **polar** – describes a molecule in which the positive and negative charges are separated
 - **polar-covalent bond** – a covalent bond in which a pair of electrons shared by two atoms is held more closely by one atom
- **Chapter Highlights**
 - Most atoms are chemically bonded to other atoms. The three major types of chemical bonding are ionic, covalent, and metallic.
 - In general, atoms of metals bond ionically with atoms of nonmetals, atoms of metals bond metallically with each other, and atoms of nonmetals bond covalently with each other.

6.2: COVALENT BONDING AND MOLECULAR COMPOUNDS

- **Vocabulary**
 - **molecule** – a group of atoms that are held together by chemical forces; a molecule is the smallest unit of matter that can exist by itself and retain all of a substance's chemical properties
 - **molecular compound** – a chemical compound whose simplest units are molecules
 - **chemical formula** – a combination of chemical symbols and numbers to represent a substance
 - **molecular formula** – a chemical formula that shows the number and kinds of atoms in a molecule, but not the arrangement of the atoms
 - **electron-dot notation** – an electron configuration notation in which only the valence electrons of an atom of a particular element are shown, indicated by dots placed around the element's symbol
 - **Lewis structure** – a structural formula in which electrons are represented by dots; dot pairs or dashes between two atomic symbols represent pairs in covalent bonds.
See examples at right for LEWIS STRUCTURES of NF₃ and OCl₂
 - **structural formula** – a formula that indicates the location of the atoms, groups, or ions relative to one another in a molecule and that indicates the number and location of chemical bonds
 - **single bond** – a covalent bond in which two atoms share one pair of electrons
 - **multiple bond** – a bond in which the atoms share more than one pair of electrons, such as a double bond or a triple bond
 - **resonance** – the bonding in molecules or ions that cannot be correctly represented by a single Lewis structure
- **Chapter Highlights**
 - Atoms in molecules are joined by covalent bonds. In a covalent bond, two atoms share one or more pairs of electrons.



Holt Chemistry Chapter 6 Test

Neil D. Jespersen, Pamela Kerrigan



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Chemistry David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom Laboratory Experiments to Accompany "Modern Chemistry," Fredus Nelson Peters, 1902 Lab Experiments Modern Chemistry Harold Clark Metcalfe, 1982 **Modern Methods of Drug Discovery** Alexander Hillisch, Rolf Hilgenfeld, 2012-11-28 Research in the pharmaceutical industry today is in many respects quite different from what it used to be only fifteen years ago There have been dramatic changes in approaches for identifying new chemical entities with a desired biological activity While chemical modification of existing leads was the most important approach in the 1970s and 1980s high throughput screening and structure based design are now major players among a multitude of methods used in drug discovery Quite often companies favor one of these relatively new approaches over the other e g screening over rational design or vice versa but we believe that an intelligent and concerted use of several or all methods currently available to drug discovery will be more successful in the medium term What has changed most significantly in the past few years is the time available for identifying new chemical entities Because of the high costs of drug discovery projects pressure for maximum success in the shortest possible time is higher than ever In addition the multidisciplinary character of the field is much more pronounced today than it used to be As a consequence researchers and project managers in the pharmaceutical industry should have a solid knowledge of the more important methods available to drug discovery because it is the rapidly and intelligently combined use of these which will determine the success or failure of preclinical projects **Essentials of Modern Chemistry** Charles Elwood Dull, 1918

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1. Understanding the eBook Holt Chemistry Chapter 6 Test
 - The Rise of Digital Reading Holt Chemistry Chapter 6 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Chapter 6 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Chapter 6 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Chapter 6 Test
 - Personalized Recommendations
 - Holt Chemistry Chapter 6 Test User Reviews and Ratings
 - Holt Chemistry Chapter 6 Test and Bestseller Lists
5. Accessing Holt Chemistry Chapter 6 Test Free and Paid eBooks
 - Holt Chemistry Chapter 6 Test Public Domain eBooks
 - Holt Chemistry Chapter 6 Test eBook Subscription Services
 - Holt Chemistry Chapter 6 Test Budget-Friendly Options

6. Navigating Holt Chemistry Chapter 6 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Chapter 6 Test Compatibility with Devices
 - Holt Chemistry Chapter 6 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Chapter 6 Test
 - Highlighting and Note-Taking Holt Chemistry Chapter 6 Test
 - Interactive Elements Holt Chemistry Chapter 6 Test
8. Staying Engaged with Holt Chemistry Chapter 6 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Chapter 6 Test
9. Balancing eBooks and Physical Books Holt Chemistry Chapter 6 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Chapter 6 Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Chemistry Chapter 6 Test
 - Setting Reading Goals Holt Chemistry Chapter 6 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Chemistry Chapter 6 Test
 - Fact-Checking eBook Content of Holt Chemistry Chapter 6 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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