



## GRADE 10 LIFE SCIENCES MARCH TEST

### SECTION A

#### QUESTION 1

- 1.1 Various options are provided as possible answers to the following questions. Choose the correct answer and write **ONLY** the **LETTER** (A-D) next to the **QUESTION NUMBER** (1.1.1-1.1.4) from your notebook. Example: 1.1.1-2.

On 10 January conducted an investigation over a period of time (only afternoon at times) to determine the wind speed at various altitudes.

Answer question 1.1.1-1.1.4 using the above-mentioned statement.

- 1.1.1 Which of the following options would be the independent variable?
- Wind speed
  - Altitude
  - Time of day
  - Location of the
- 1.1.2 Which of the following options would be considered the dependent variable?
- Wind speed
  - Altitude
  - Time of day
  - Location of the
- 1.1.3 Which of the following options would be a fixed variable?
- Wind speed
  - Altitude
  - Time of day
  - Location of the
- 1.1.4 The following is **NOT** an abiotic factor in the ecosystem of a grasshopper.
- The climate
  - The grasshopper's food source
  - The grasshopper's habitat
  - The grasshopper's colour

- 1.1.5 Which of the following is the most appropriate example of a population?

- All organisms in a pond
- All the grasshoppers in a field
- All the grasshoppers in a field
- All the grasshoppers in a field

[2-10]

- 1.2 Give the correct **BIOLOGICAL TERM** for each of the following statements. Write **ONLY** the **TERM** next to the **QUESTION NUMBER** (1.2.1-1.2.4).

- 1.2.1 A group of individuals of the same species living in the same area.
- 1.2.2 A group of individuals of the same species living in the same area.
- 1.2.3 A group of individuals of the same species living in the same area.
- 1.2.4 A group of individuals of the same species living in the same area.

[2-10]

- 1.3 Indicate whether each of the statements in **COLUMN I** applies to **A ONLY**, **B ONLY**, **BOTH A and B** or **NONE** of the items in **COLUMN II**. Write **A ONLY**, **B ONLY**, **BOTH A and B** or **NONE** next to the **QUESTION NUMBER** (1.3.1-1.3.3).

	COLUMN I	COLUMN II
1.3.1	Location of the Savanna biome in South Africa	A. Climate B. Soil C. Water D. Sunlight
1.3.2	Climate refers to	A. The weather B. The temperature C. The humidity D. The wind speed
1.3.3	Stated that the wind speed affects the temperature in the atmosphere.	A. True B. False

[2-10]  
TOTAL SECTION A = 20

# Grd10 Memo Life Science Practice P2

**RD Boyd**



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## **Table of Contents Grd10 Memo Life Science Practice P2**

1. Understanding the eBook Grd10 Memo Life Science Practice P2
  - The Rise of Digital Reading Grd10 Memo Life Science Practice P2
  - Advantages of eBooks Over Traditional Books
2. Identifying Grd10 Memo Life Science Practice P2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grd10 Memo Life Science Practice P2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grd10 Memo Life Science Practice P2
  - Personalized Recommendations
  - Grd10 Memo Life Science Practice P2 User Reviews and Ratings

- Grd10 Memo Life Science Practice P2 and Bestseller Lists
- 5. Accessing Grd10 Memo Life Science Practice P2 Free and Paid eBooks
  - Grd10 Memo Life Science Practice P2 Public Domain eBooks
  - Grd10 Memo Life Science Practice P2 eBook Subscription Services
  - Grd10 Memo Life Science Practice P2 Budget-Friendly Options
- 6. Navigating Grd10 Memo Life Science Practice P2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Grd10 Memo Life Science Practice P2 Compatibility with Devices
  - Grd10 Memo Life Science Practice P2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grd10 Memo Life Science Practice P2
  - Highlighting and Note-Taking Grd10 Memo Life Science Practice P2
  - Interactive Elements Grd10 Memo Life Science Practice P2
- 8. Staying Engaged with Grd10 Memo Life Science Practice P2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grd10 Memo Life Science Practice P2
- 9. Balancing eBooks and Physical Books Grd10 Memo Life Science Practice P2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grd10 Memo Life Science Practice P2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grd10 Memo Life Science Practice P2
  - Setting Reading Goals Grd10 Memo Life Science Practice P2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grd10 Memo Life Science Practice P2
  - Fact-Checking eBook Content of Grd10 Memo Life Science Practice P2
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2 + \text{aq} \rightarrow \text{NH}_3 + \text{aq}$  2 practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^{-}$   $\text{Br}_2 + \text{l} \rightarrow \text{Br}^{-} + \text{aq}$  previous

topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{H}_2\text{SO}_3 \rightarrow \text{Mn}^{2+} + \text{SO}_4^{2-}$  what will be the stoichiometric coefficient of  $\text{H}_2\text{SO}_3$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web may 25 2021 balancing redox equations using half reactions another way to balance redox reactions is by the half reaction method this technique involves breaking an equation into its two separate components the oxidation reaction and

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O}_2$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O}_2$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O}_2 \rightarrow 2\text{OH}^- + 2\text{H}_2\text{O}$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h<sub>2</sub>o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh<sup>-</sup> ions or the oh

$\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between  $\text{ClO}^-$  and  $\text{Cr}(\text{OH})_4^-$  in basic solution created by jay

### **balancing redox reactions in basic solution problems 1 10** - Aug 06 2023

web bonus problem here is a balanced net ionic equation  $3\text{Br}_2 + 6\text{OH}^- \rightarrow 3\text{Br}_2 + 3\text{H}_2\text{O}$  the above equation takes place in two stages first bromine reacts with  $\text{OH}^-$  ions to form bromide ions and  $\text{BrO}^-$  ions second the  $\text{BrO}^-$  ions react to form bromide ions and bromate  $\text{BrO}_3^-$  ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

### balancing redox reactions in acidic solution problems 1 10 - Sep 07 2023

web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

### **worked example balancing a redox equation in acidic solution** - Mar 21 2022

web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

### *worksheet 5 balancing redox reactions in acid and basic* - Aug 26 2022

web balance each half reaction in basic solution  $4\text{SO}_3^{2-} + 2\text{MnO}_2 + \text{Cr}_3\text{O}_3 + \text{NO}_3^- \rightarrow \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+} + 10\text{H}_2\text{O}$   $11\text{PbO}_2 + \text{IO}_3^- + 2\text{H}^+ \rightarrow \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{SO}_4$

### **balancing redox equations practice exercises enig periodic** - Oct 08 2023

web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$  and  $\text{OH}^-$  in this order balance the atoms in the equation apart from o and h

**redox reactions questions practice khan academy** - Apr 21 2022

web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$  choose 1 answer  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{C}_6\text{H}_6\text{O}_6$  c  $\text{HNO}_2$  and  $\text{C}_6\text{H}_6\text{O}_6$  d  $\text{HNO}_2$  and  $\text{NO}_2$