

Objective type

- Q.1 Which of the following is not a general characteristic of equilibria involving physical processes?
- (1) Equilibrium is possible only in a closed system at a given temperature.
 - (2) All measurable properties of the system remain constant.
 - (3) All the physical processes stop at equilibrium.
 - (4) The opposing processes occur at the same rate and there is dynamic but stable condition.
- Q.2 Which of the following statements is incorrect?
- (1) In equilibrium mixture of ice and water kept in perfectly insulated flask mass of ice and water does not change with time.
 - (2) The intensity of red colour increases when oxalic acid is added to a solution containing iron (III) nitrate and potassium thiocyanate.
 - (3) On addition of catalyst the equilibrium constant value is not affected.
 - (4) Equilibrium constant for a reaction with negative ΔH value decreases as the temperature increases.
- Q.3 When hydrochloric acid is added to cobalt nitrate solution at room temperature, the following reaction takes place and the reaction mixture becomes blue. On cooling the mixture it becomes pink. On the basis of this information mark the correct answer.
- $$\underset{\text{(pink)}}{[\text{Co}(\text{H}_2\text{O})_6]^{3+}}(\text{aq}) + 4\text{Cl}^{-}(\text{aq}) \rightleftharpoons \underset{\text{(blue)}}{[\text{CoCl}_4]^{2-}}(\text{aq}) + 6\text{H}_2\text{O}(\text{l})$$
- (1) $\Delta H > 0$ for the reaction
 - (2) $\Delta H < 0$ for the reaction
 - (3) $\Delta H = 0$ for the reaction
 - (4) The sign of ΔH cannot be predicted on the basis of this information.
- Q.4 What will be the correct order of vapour pressure of water, acetone and ether at 30°C . Given that among these compounds, water has maximum boiling point and ether has minimum boiling point?
- (1) Water < ether < acetone
 - (2) Water < acetone < ether
 - (3) Ether < acetone < water
 - (4) Acetone < ether < water
- Q.5 In which of the following reactions, the equilibrium remains unaffected on addition of small amount of argon at constant volume?
- (1) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 - (2) $\text{PCl}_5(\text{g}) \rightleftharpoons \text{PCl}_3(\text{g}) + \text{Cl}_2(\text{g})$
 - (3) $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
 - (4) The equilibrium will remain unaffected in all the three cases

Assertion & Reason

- Q.6 Assertion (A): In the dissociation of PCl_5 at constant pressure and temperature addition of helium at equilibrium increases the dissociation of PCl_5 .
- Reason (R) : Helium removes Cl_2 from the field of action.
- (1) Both A and R are true and R is correct explanation of A.
 - (2) Both A and R are true but R is not correct explanation of A.
 - (3) A is true but R is false.
 - (4) Both A and R are false.

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Table of Contents Mcqs Of Environmental Chemistry

1. Understanding the eBook Mcqs Of Environmental Chemistry
 - The Rise of Digital Reading Mcqs Of Environmental Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcqs Of Environmental Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcqs Of Environmental Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcqs Of Environmental Chemistry
 - Personalized Recommendations

- Mcqs Of Environmental Chemistry User Reviews and Ratings
- Mcqs Of Environmental Chemistry and Bestseller Lists
- 5. Accessing Mcqs Of Environmental Chemistry Free and Paid eBooks
 - Mcqs Of Environmental Chemistry Public Domain eBooks
 - Mcqs Of Environmental Chemistry eBook Subscription Services
 - Mcqs Of Environmental Chemistry Budget-Friendly Options
- 6. Navigating Mcqs Of Environmental Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Mcqs Of Environmental Chemistry Compatibility with Devices
 - Mcqs Of Environmental Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcqs Of Environmental Chemistry
 - Highlighting and Note-Taking Mcqs Of Environmental Chemistry
 - Interactive Elements Mcqs Of Environmental Chemistry
- 8. Staying Engaged with Mcqs Of Environmental Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcqs Of Environmental Chemistry
- 9. Balancing eBooks and Physical Books Mcqs Of Environmental Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcqs Of Environmental Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcqs Of Environmental Chemistry
 - Setting Reading Goals Mcqs Of Environmental Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcqs Of Environmental Chemistry
 - Fact-Checking eBook Content of Mcqs Of Environmental Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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