

## HOLT MODERN CHEMISTRY -- CHAPTER 3 REVIEW

*The following pages contain the bulk (but not all) of the information for the chapter 3 test. Focus on this content, but make sure to review class notes, activities, handouts, questions, etc. If you study this document and NOTHING else, you should at least be able to PASS the test. \*\*\*\*\* Test items will be recall, examples, and/or application of this content. \*\*\*\*\**

### OUTCOMES

- 3.1: Explain law of conservation of mass (*T & R*)
- 3.1-3.2: Explain how the atom changed over time (Rutherford - Bohr models) (*F & PK*)
- 3.3: Contrast the nature and properties of the three subatomic particles: proton, neutron, and electron (*T & R*)
- 3.3: Use atomic number and mass number to find the number of subatomic particles of any given atom (*F & PK*)
- 3.3: Draw contrasting Bohr models of isotopes (*F & PK*)
- 3.3: Define mole and Avogadro's number (*T & R*)

### 3.1 THE ATOM: FROM PHILOSOPHICAL IDEA TO SCIENTIFIC THEORY

- **Vocabulary**
  - **law of conservation of mass** -- the law that states that mass cannot be created or destroyed in ordinary chemical and physical changes
  - **law of definite proportions** -- the law that states that a chemical compound always contains the same elements in exactly the same proportions by weight or mass
  - **law of multiple proportions** -- the law that states that when two elements combine to form two or more compounds, the mass of one element that combines with a given mass of the other is in the ratio of small whole numbers
- **Chapter Highlights**
  - The idea of atoms has been around since the time of the ancient Greeks. In the nineteenth century, John Dalton proposed a scientific theory of atoms that can still be used to explain properties of most chemicals today.
  - Matter and its mass cannot be created or destroyed in chemical reactions.
  - The mass ratios of the elements that make up a given compound are always the same, regardless of how much of the compound there is or how it was formed.
  - If two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element can be expressed as a ratio of small whole numbers.

### 3.2 THE STRUCTURE OF THE ATOM

- **Vocabulary**
  - **atom** -- the smallest unit of an element that maintains the chemical properties of that element
  - **nuclear forces** -- the interaction that binds protons and neutrons, protons and protons, and neutrons and neutrons together in a nucleus
- **Chapter Highlights**
  - Cathode-ray tubes supplied evidence of the existence of electrons, which are negatively charged subatomic particles that have relatively little mass.
  - Rutherford found evidence for the existence of the atomic nucleus by bombarding gold foil with a beam of positively charged particles.
  - Atomic nuclei are composed of protons, which have an electric charge of +1, and (in all but one case) neutrons, which have no electric charge.
  - Atomic nuclei have radii of about 0.001 pm (pm = picometers;  $1 \text{ pm} \times 10^{-12} \text{ m}$ ), and atoms have radii of about  $40^{-270} \text{ pm}$ .

# Holt Modern Chemistry Notes Chapter 18

**Steven M. Brown**



## Holt Modern Chemistry Notes Chapter 18:

**The Development of Modern Chemistry** Aaron J. Ihde, 1984-01-01 From ancient Greek theory to the explosive discoveries of the 20th century this authoritative history shows how major chemists their discoveries and political economic and social developments transformed chemistry into a modern science 209 illustrations 14 tables Bibliographies Indices Appendices New York Teachers' Monographs ,1915 **Teachers' Monographs** ,1915 **Reclaiming Our Legacy** Steven M. Brown, 1986 *Magick, Mayhem, and Mavericks* Cathy Cobb, 2010-09-09 Science popularizer Cathy Cobb takes a unique approach to explaining the concepts of physical chemistry by telling the story of the geniuses and eccentrics who made groundbreaking discoveries in this fascinating field that bridges chemistry physics and mathematics The result is entertaining and illuminating Her tale is about the colorful varieties of human character as well as the struggles to understand the workings of the material world Through true stories of rebels recluses heroes and rogues she helps the reader to discover how one idea built upon another and how an elegant discipline arose out of centuries of difficult trial and error Starting with the ancient Greeks Cobb takes the reader on a sweeping tour of history She shows how an understanding of basic chemical properties gradually arose out of ancient Greeks mathematics Muslim science medieval magick and the healing arts Her tour continues through the scientific revolution the emergence of physical chemistry as an independent discipline and up to the present Today physical chemists contribute to the fields of chemical physiology chemical oscillations and waves quantum mechanics and the curious and promising field of nanotechnology This absorbing eloquently written history of science is loaded with intuitive imagery everyday analogies and a colorful cast of characters who are guaranteed to entertain as well as edify The Athenaeum ,1907 **Catalog of Copyright Entries. Third Series** Library of Congress. Copyright Office, 1962 Includes Part 1 Number 2 Books and Pamphlets Including Serials and Contributions to Periodicals July December *Book Chat* ,1888 The academy ,1877 **Statistical Methods in Quantum Optics 1** Howard J. Carmichael, 2013-04-17 As a graduate student working in quantum optics I encountered the question that might be taken as the theme of this book The question definitely arose at that time though it was not yet very clearly defined there was simply some deep irritation caused by the work I was doing something quite fundamental I did not understand Of course so many things are not understood when one is a graduate student However my nagging question was not a technical issue not merely a mathematical concept that was difficult to grasp It was a sense that certain elementary notions that are accepted as starting points for work in quantum optics somehow had no fundamental foundation no identifiable root My inclination was to mine physics vertically and here was a subject whose tunnels were dug horizontally There were branches certainly going up and going down Nonetheless something major in the downwards direction was missing at least in my understanding no doubt others understood the connections downwards very well In retrospect I can identify the irritation Quantum optics deals primarily with dynamics quantum dynamics and in doing so makes extensive use of words like quantum fluctuations and

quantum noise The words seem harmless enough Surely the ideas behind them are quite clear after all quantum mechanics is a statistical theory and in its dynamical aspects it is therefore a theory of fluctuations But there was my problem Nothing in Schrodinger s equation fluctuates

## Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Holt Modern Chemistry Notes Chapter 18**

In a world inundated with monitors and the cacophony of instant communication, the profound power and mental resonance of verbal artistry frequently diminish into obscurity, eclipsed by the constant assault of sound and distractions. However, located within the musical pages of **Holt Modern Chemistry Notes Chapter 18**, a captivating function of fictional elegance that pulses with raw feelings, lies an unique trip waiting to be embarked upon. Written by way of a virtuoso wordsmith, this mesmerizing opus manuals viewers on a mental odyssey, lightly exposing the latent potential and profound impact stuck within the delicate web of language. Within the heart-wrenching expanse of the evocative examination, we shall embark upon an introspective exploration of the book is central styles, dissect its captivating writing style, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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