

## Concept

For a gas phase reaction



The equilibrium constant is written as

$$K = \left[ \frac{(P_C)^c (P_D)^d}{(P_A)^a (P_B)^b} \right]_{eq}$$

**Note:** Remember that solids and liquids are not included in the equilibrium constant expression.

## Solution



$$K = \left[ \frac{(P_{\text{H}_2\text{S}})^8}{(P_{\text{H}_2})^8} \right]_{eq}$$



$$K = \left[ \frac{(P_{\text{COCl}_2}) (P_{\text{H}_2})}{(P_{\text{Cl}_2})} \right]$$



$$K = [P_{\text{CO}_2}]_{eq}$$



$$K = \left[ \frac{1}{(P_{\text{C}_2\text{H}_2})^3} \right]_{eq}$$

# Holt Modern Chemistry Chapter 9 Answers

**Christian G. Meyer**



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