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web we can write a balanced nuclear equation for this reaction as let s do a practice example identify the product x by writing a balanced nuclear equation $^{137}_{55}\text{Cs} \rightarrow ^{137}_{56}\text{Ba} + x$ because the mass number is the same on both sides of the equation x must have a

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web identify the missing particle in the following nuclear reaction $^{1}_{0}\text{n} + ^{235}_{92}\text{U} \rightarrow ^{2}_{1}\text{n} + ^{137}_{52}\text{Te} + ^{97}_{38}\text{Sr} + ^{97}_{40}\text{Zr}$

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web answer 4 correct in the reactants and products must be the same part b what is the value of in the following nuclear reaction $^{6}_{5}\text{Li} + ^{9}_{4}\text{Be} \rightarrow ^{12}_{6}\text{C} + ^{4}_{2}\text{He}$ express your answer as an integer hint 1 how to approach the problem use rule 2 the total of the subscripts

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